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Aqueous Equilibria: Part 1 - Buffers & Titrations

PSI AP Chemistry

Name _____

Multiple Choice questions

Common-Ion Effect

1) The pH of a solution that contains 0.818 M acetic acid ($K_a = 1.77 \times 10^{-5}$) and 0.172 M sodium acetate is _____.

- A) 4.077 B) 5.434 C) 6.571 D) 8.370 E) 9.922

2) Calculate the pH of a solution prepared by dissolving 0.37 mol of formic acid (HCO_2H) and 0.23 mol of sodium formate (NaCO_2H) in water sufficient to yield 1.00 L of solution. The K_a of formic acid is 1.77×10^{-4} .

- A) 2.1 B) 10.46 C) 3.55 D) 2.31 E) 3.95

3) Calculate the pH of a solution prepared by dissolving 0.25 mol of benzoic acid ($\text{C}_6\text{H}_5\text{O}_2\text{H}$) and 0.15 of sodium benzoate ($\text{NaC}_6\text{H}_5\text{O}_2$) in water sufficient to yield 1.00 L of solution. The K_a of benzoic acid is 6.50×10^{-5} .

- A) 4.41 B) 2.40 C) 3.97 D) 10.08 E) 4.19

4) Calculate the pH of a solution prepared by dissolving 0.15 mol of benzoic acid (HBz) and 0.30 mol of sodium benzoate in water sufficient to yield 1.00L of solution. The K_a of benzoic acid is 6.50×10^{-5} .

- A) 2.52 B) 3.89 C) 4.49 D) 10.16 E) 4.20

5) Calculate the pH of a solution prepared by dissolving 0.750 mol of NH_3 and 0.250 mol of NH_4Cl in water sufficient to yield 1.00 L of solution. The K_b of ammonia is 1.77×10^{-5} .

- A) 5.22 B) 4.27 C) 9.73 D) 8.78 E) 0.89

6) The pH of a solution prepared by dissolving 0.350 mol of solid methylamine hydrochloride ($\text{CH}_3\text{NH}_3\text{Cl}$) in 1.00L of 1.10M methylamine (CH_3NH_2) is _____. The K_b for methylamine is 4.4×10^{-4} . (Assume the final volume is 1.00 L.)

- A) 1.86 B) 2.66 C) 10.15 D) 11.14 E) 10.64

Buffers

7) What does a buffer do?

- a) Keeps the pH of a solution constant.

Practice Multiple Choice For Chem Aqueous Equilibria

Wenbin Ji



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1. Understanding the eBook Practice Multiple Choice For Chem Aqueous Equilibria
 - The Rise of Digital Reading Practice Multiple Choice For Chem Aqueous Equilibria
 - Advantages of eBooks Over Traditional Books
2. Identifying Practice Multiple Choice For Chem Aqueous Equilibria
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Practice Multiple Choice For Chem Aqueous Equilibria
 - User-Friendly Interface
4. Exploring eBook Recommendations from Practice Multiple Choice For Chem Aqueous Equilibria
 - Personalized Recommendations

- Practice Multiple Choice For Chem Aqueous Equilibria User Reviews and Ratings
- Practice Multiple Choice For Chem Aqueous Equilibria and Bestseller Lists
- 5. Accessing Practice Multiple Choice For Chem Aqueous Equilibria Free and Paid eBooks
 - Practice Multiple Choice For Chem Aqueous Equilibria Public Domain eBooks
 - Practice Multiple Choice For Chem Aqueous Equilibria eBook Subscription Services
 - Practice Multiple Choice For Chem Aqueous Equilibria Budget-Friendly Options
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 - Practice Multiple Choice For Chem Aqueous Equilibria Enhanced eBook Features
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 - Highlighting and Note-Taking Practice Multiple Choice For Chem Aqueous Equilibria
 - Interactive Elements Practice Multiple Choice For Chem Aqueous Equilibria
- 8. Staying Engaged with Practice Multiple Choice For Chem Aqueous Equilibria
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Practice Multiple Choice For Chem Aqueous Equilibria
- 9. Balancing eBooks and Physical Books Practice Multiple Choice For Chem Aqueous Equilibria
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Practice Multiple Choice For Chem Aqueous Equilibria
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Practice Multiple Choice For Chem Aqueous Equilibria
 - Setting Reading Goals Practice Multiple Choice For Chem Aqueous Equilibria
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Practice Multiple Choice For Chem Aqueous Equilibria
 - Fact-Checking eBook Content of Practice Multiple Choice For Chem Aqueous Equilibria

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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