



Republic of the Philippines
Department of Education
REGION III

REGIONAL DIAGNOSTIC ASSESSMENT IN PHYSICAL SCIENCE

Name _____ Score _____

Direction: Choose the correct answer for each item.

1. What type of nuclear synthesis is believed to be the origin of the production of heavy elements?
A. Big Bang Nucleosynthesis
B. Primordial Nucleosynthesis
C. Stellar Nucleosynthesis
D. Supernova Nucleosynthesis
2. What is the ability of an atom in a compound to attract electrons towards itself called?
A. Electromotive force
B. Electron Affinity
C. Electronegativity
D. Ionization energy
3. Which of the following is the measure of a liquid's resistance to flow?
A. Boiling point
B. Melting point
C. Solubility
D. Viscosity
4. Which of the following is the building block of protein?
A. Amino acid
B. Glycogen
C. Monomer
D. Nucleotide
5. What do you call the substance that increases the rate of chemical reaction without itself being consumed by the reaction?
A. Activation energy
B. Catalyst
C. Product
D. Reactant
6. Which of the following describes a stellar nucleosynthesis?
A. It is the process by which elements are formed within stars.
B. It is the formation of elements during a supernova explosion
C. It is the formation of light elements such as helium and hydrogen
D. It is the process by which elements are produced within gas clouds
7. Which of the following statements is TRUE about non polar molecules?
A. They have high boiling point
B. They have high melting point
C. They have low surface tension
D. They have low vapor pressure
8. Which will result in an ion- dipole interaction?
A. Attraction between an ion and a polar molecule
B. Attraction between a polar with a polar molecule
C. Repulsion between dipole and another dipole
D. Repulsion between a polar and a nonpolar molecule
9. Why are lipids insoluble in water?
A. Because Lipids are hydrophilic
B. Because Lipids are hydrophobic
C. Because Lipids are neutral
D. Because Lipids are ionic compound
10. What will be the effect of temperature rise in a chemical reaction?
A. It increases molecular concentration and slows down reaction rate
B. It increases the number of effective collisions among reactants, thus increasing the reaction rate
C. It decreases the system's energy and limits how much further reaction can take place
D. It decreases the number of collisions and the energy of molecules to limit further reaction

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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \quad \Delta H = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta H = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g}) \quad \Delta H = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g})$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l}) \quad \Delta H = 1411 \text{ kJ/mol}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l}) \quad \Delta H = 1560 \text{ kJ/mol}$
 $\text{H}_2\text{O}(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta H = 285 \text{ kJ/mol}$

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web sep 22 2021 hess's law states that the enthalpy change of an overall process is equal to the sum of the enthalpy changes of its individual steps hess's law example 12.1 determine ΔH for the target reaction $\text{C}_2\text{H}_4(\text{g}) + 2 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l})$ given the following information

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web libretexts states hess's law states that regardless of the multiple stages or steps of a reaction the total enthalpy change for the reaction is the sum of all changes libretexts quote hess's law quote 2019 which determines the enthalpy change value hess's law is written as $\Delta H_{\text{target}} = \Delta H_1 + \Delta H_2 + \Delta H_3$ or $\Delta H_{\text{target}} = \sum \Delta H_{\text{known}}$

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web hess's law states that if a process can be expressed as the sum of two or more steps the enthalpy change for the overall

process is the sum of the Δh values for each step to use hess's law two principles must be understood one if an equation is reversed the sign of the Δh value is also reversed two if an equation is multiplied by a

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web hess's law states that the total enthalpy change for the reaction will be the sum of all those changes no matter how many different steps or stages in the reaction there are cohen 2016 the equations for the reactions in the experiment done are as follows $\text{NaOH} + \text{Na} \rightarrow \text{NaOH}$

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web hess's law worksheet answers 1 calculate h for the reaction $\text{C}_2\text{H}_4 + \text{H}_2 \rightarrow \text{C}_2\text{H}_6$ from the following data $\text{C}_2\text{H}_4 + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$ $\Delta h = 1411 \text{ kJ}$ $\text{C}_2\text{H}_6 + \frac{3}{2}\text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2\text{O}$ $\Delta h = 1560 \text{ kJ}$

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web mar 13 2023 be able to use hess's law to calculate reaction enthalpies first law of thermodynamics energy is the ability to do work or transfer heat work is the transfer of energy from one body to another in a sense work is energy in the process of transfer

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web may 8 2015 hess law lab by maya parks partners ben seufert kelsea floyd convert joules to kj in your final answer multiply the mass by the change in temperature and the c given trial 1 use hess s law and the following equations and Δh values to determine the heat of reaction for the reaction c 2 h 4

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web thus for reaction 9 2 in which 1 mole of H_2O and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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web hess s law enthalpy of reaction and hess s law abstract the goal of this lab is to perform experiments and practice hess s law by determining enthalpy change in chemical reactions in this lab three experiments were conducted in order to test hess s law the experiments required observing and collecting data on three different reactions

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions 1 $\text{NaOH (s)} + \text{NaOH (aq)} \rightarrow \text{H}_2\text{O (l)}$ 2 $\text{NaOH (s)} + \text{HCl (aq)} \rightarrow \text{H}_2\text{O (l)}$ 3 $\text{NaOH (aq)} + \text{HCl (aq)} \rightarrow \text{H}_2\text{O (l)}$