# EXAMINATION DATA SHEET FOR THE PHYSICAL SCIENCES (CHEMISTRY)

#### TABLE 1 PHYSICAL CONSTANTS

NAME	SYMBOL	VALUE
Magnitude of charge on electron	e	1,6 × 10 <sup>-19</sup> C
Mass of an electron	m <sub>e</sub>	9,1 × 10 <sup>-31</sup> kg
Standard pressure	p*	1,01 × 10 <sup>5</sup> Pa
Molar gas volume at STP	V <sub>m</sub>	22,4 dm <sup>3</sup> ·mol <sup>-1</sup>
Standard temperature	T*	273 K
Avogadro's constant	N <sub>A</sub>	6,02 × 10 <sup>23</sup> mol <sup>-1</sup>
Faraday's constant	F	96 500 C · mol-1

#### TABLE 2 CHEMISTRY FORMULAE

$$n = \frac{m}{M}$$

$$n = \frac{N}{N_A}$$

$$n = \frac{V}{V_m}$$

$$G = \frac{n}{V} \text{ OR } G = \frac{m}{MV}$$

$$K_w = \begin{bmatrix} H_3O^+ \end{bmatrix} \cdot \begin{bmatrix} OH^- \end{bmatrix} = 1 \times 10^{-14}$$

$$\text{at 25 °C (298 K)}$$

$$Q = It$$

$$Q = nF$$

$$E_{cell}^\theta = E_{oxidising agent}^\theta - E_{reducing agent}^\theta$$

$$E_{cell}^\theta = E_{oxidising agent}^\theta - E_{reducing agent}^\theta$$

# Physical Science P2 Grade1exampler

Judd E. Hollander

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