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4.14.

$$\left( \frac{1000 \text{ g}}{1000 \text{ g}} \right) \left( \frac{1000 \text{ g}}{1000 \text{ g}} \right) = 1.0 \text{ g/g}$$

$$\text{mass of H}_2\text{O} = 1.0 \text{ g/g} \times 1.0 \text{ g} = 1.0 \text{ g} \left( \frac{1000 \text{ g}}{1000 \text{ g}} \right) = 1.0 \text{ g}$$

4.15. (a) The water molecule has two lone pairs of electrons.

(b) The water molecule has two lone pairs of electrons.

(c) The water molecule has two lone pairs of electrons.

4.16.

(a) The water molecule has two lone pairs of electrons.

$$\text{mass of H}_2\text{O} = 1.0 \text{ g/g} \times 1.0 \text{ g} = 1.0 \text{ g} \left( \frac{1000 \text{ g}}{1000 \text{ g}} \right) = 1.0 \text{ g}$$

4.17.

(a) The water molecule has two lone pairs of electrons.

$$\text{mass of H}_2\text{O} = 1.0 \text{ g/g} \times 1.0 \text{ g} = 1.0 \text{ g} \left( \frac{1000 \text{ g}}{1000 \text{ g}} \right) = 1.0 \text{ g}$$

(b) The water molecule has two lone pairs of electrons.

$$\text{mass of H}_2\text{O} = 1.0 \text{ g/g} \times 1.0 \text{ g} = 1.0 \text{ g} \left( \frac{1000 \text{ g}}{1000 \text{ g}} \right) = 1.0 \text{ g}$$

4.18.

(a) The water molecule has two lone pairs of electrons.

$$\text{mass of H}_2\text{O} = 1.0 \text{ g/g} \times 1.0 \text{ g} = 1.0 \text{ g} \left( \frac{1000 \text{ g}}{1000 \text{ g}} \right) = 1.0 \text{ g}$$

# Physical Chemistry Ball Solution Manual

**Richard Bailey**



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