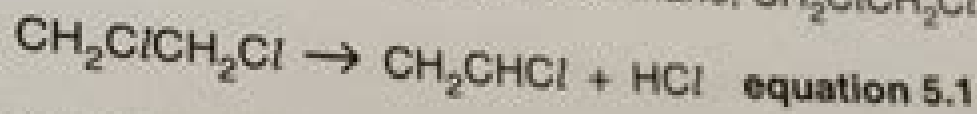


Chloroethene, CH_2CHCl , is made from 1,2-dichloroethane, $\text{CH}_2\text{ClCH}_2\text{Cl}$.



- (a) A chemical plant uses 19.80 tonnes of 1,2-dichloroethane to make 11.25 tonnes of chloroethene.

Calculate the percentage yield of chloroethene.

(1 tonne = $1.00 \times 10^6 \text{ g}$)

$$\% = \frac{\text{actual amount in mol}}{\text{Theoretical amount in mol}} \times 100$$

$$= \frac{\frac{11.25 \times 10^6}{62.5}}{\frac{19.8 \times 10^6}{98}} \times 100 =$$

Ocr Chemistry Past Paper F322 June 2014

Judd E. Hollander



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