



## atomic radius-practice problems

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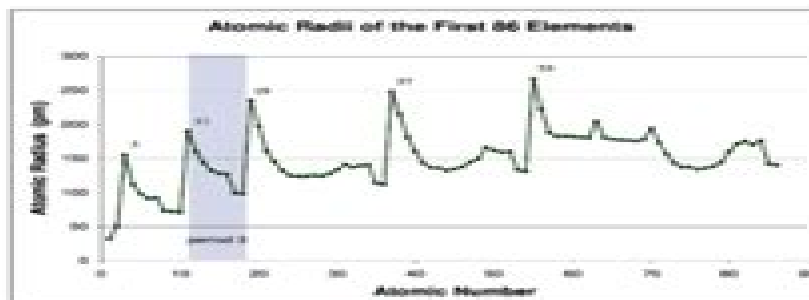
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### trends on the periodic table – atomic radius

The **atomic radius** describes the distance between the nucleus and the outermost electron. This distance is very small so it is often reported in **picometers** ( $1 \times 10^{-12}$  m) or **angstroms** ( $1 \times 10^{-10}$  m). One way of measuring the atomic radius is to observe the distance between two bonding atoms. The distance between the nuclei of the bonding atoms is equal to twice the atomic radius. This distance is called the **bonding atomic radius** and is shorter than the non-bonded radius.



A pattern is observed when the atomic radius is plotted versus the atomic number, as seen above. As might be expected the largest atoms are found at the bottom of the periodic table, however, as you go across a period, the atomic radius decreases (see highlighted region in the graph above). All of the elements in a period have the same orbitals which have the same energy and same size. Additionally, the elements that lay on the right side of the periodic table have nuclei which contain more protons. This results in a larger attraction of the electrons and a smaller radius. This explains why argon (atomic number = 18) has a smaller radius than sodium (atomic number = 11).

**Directions:** Answer the following questions on a separate piece of paper.

1. Estimate the size of the smallest atom in the chart above. Estimate the size of the largest.
2. Which five elements are represented by the tallest 'peaks' in the graph above?
3. Write the approximate radius for each of the noble gases in the table below.

element	atomic radius	element	atomic radius	element	atomic radius
helium		argon		xenon	
neon		krypton		radon	

4. What is the general trend for atomic radius as you go down the noble gas family?
5. What is the general trend as you move through period 3 (from element 11 to element 18)?

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# Practice Problems With Periodic Trends

**Brendan G. Carr**



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