

Specific Heat

DIRECTIONS: Use formula $q = (m)(\Delta T)(C)$ to solve the following problems. Show all work and units. Use <http://websiteschools.com> for specific heat of various substances.

1. A 30g piece of metal absorbs 1,200 joules of heat energy and its temperature change from 25°C to 175°C. Calculate the specific capacity of the metal. What is the likely metal?
2. How many joules of heat are needed to raise the temperature of 15.0 g of iron from 11°C to 68°C, if the specific heat of iron is 0.45 J/g°C?
3. To what temperature will a 50.0 g piece of brass raise if it absorbs 6 kilojoules of heat and its specific heat capacity is 0.38 J/g°C? The initial temperature of the brass is 20.0°C.
4. Identify the substance if 1500.0 g of the substance absorbs 74,500 joules of heat, and its temperature changes from 25°C to 47°C.
5. 100.0 g of 1.0°C water is heated until its temperature is 37°C. If the specific heat of water is 4.18 J/g°C, calculate the amount of heat energy needed to cause this rise in temperature.
6. 25.0 g of sodium is heated from 30°C to 160°C, and absorbs 750 joules of heat in the process. Calculate the specific heat capacity of sodium.
7. What is the specific heat capacity of quartz glass if 200.0 g of the glass absorbs 2520 joules of heat and the temperature rises 18.0°C?
8. If a sample of titanium is initially at 25°C, what is its final temperature if 250.0 g of titanium absorbs 1.0 kilojoules of heat, and the specific heat of titanium is 0.523 J/g°C?
9. How much energy would be absorbed by 300.0 g of steel to increase its temperature from 100.0°C to 450.0°C? (C of steel = .48 J/g °C)
10. When 55.0 g of platinum drops its temperature from 300.0°C to 250.0°C, how much heat energy is released? (C of platinum = .133 J/g °C)
11. How much energy is required to heat 125.0 g of lithium from 2.0°C to 30.0°C? (C of lithium = 3.6 J/g°C)
12. How much heat (in kJ) is given out when 350.0 g of lead cools from 200.0°C to 30.0°C?

Practice Problems For Specific Heat

Joacim Rocklöv



Practice Problems For Specific Heat:

Decoding **Practice Problems For Specific Heat**: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Practice Problems For Specific Heat**," a mesmerizing literary creation penned with a celebrated wordsmith, readers set about an enlightening odyssey, unraveling the intricate significance of language and its enduring affect our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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