

**Concept-Development  
Practice Page**

**32-1**

**Coulomb's Law**

1. The diagram is of a hydrogen atom.

- Label the proton in the nucleus with a  $+$  sign and the orbital electron with a  $-$  sign.
- The electrical interaction between the nucleus and the orbital electron is a force of \_\_\_\_\_ (attraction) \_\_\_\_\_ (repulsion).
- According to Coulomb's law,

$$F = k \frac{q_1 q_2}{r^2}$$

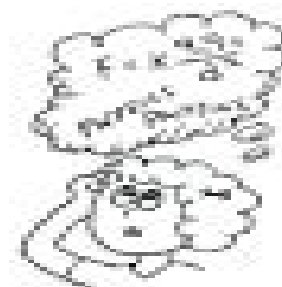
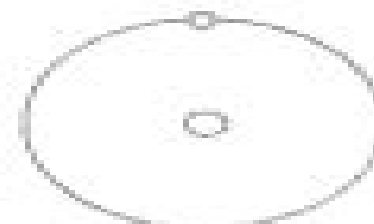
if the charge of either the nucleus or the orbital electron were greater, the force between the nucleus and the electron would be \_\_\_\_\_ (greater) \_\_\_\_\_ (less).

and if the distance between the nucleus and electron were greater the force would be \_\_\_\_\_ (greater) \_\_\_\_\_ (less).

If the distance between the nucleus and electron were doubled, the force would be \_\_\_\_\_ (1/4 as much) \_\_\_\_\_ (1/2 as much) \_\_\_\_\_ (two times as much) \_\_\_\_\_ (four times as much).

2. Consider the electric force between a pair of charged particles a certain distance apart. By Coulomb's law:

- If the charge on one of the particles is doubled, the force is \_\_\_\_\_ (unchanged) \_\_\_\_\_ (halved) \_\_\_\_\_ (doubled) \_\_\_\_\_ (quadrupled).
- If instead the charge on both particles is doubled, the force is \_\_\_\_\_ (unchanged) \_\_\_\_\_ (halved) \_\_\_\_\_ (doubled) \_\_\_\_\_ (quadrupled).
- If instead the distance between the particles is halved, the force is \_\_\_\_\_ (unchanged) \_\_\_\_\_ (halved) \_\_\_\_\_ (doubled) \_\_\_\_\_ (quadrupled).
- If the distance is halved and the charge of both particles is doubled, the force is \_\_\_\_\_ times as great.



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**VM Jensen**



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## **Table of Contents Practice Page 32 1**

1. Understanding the eBook Practice Page 32 1
  - The Rise of Digital Reading Practice Page 32 1
  - Advantages of eBooks Over Traditional Books
2. Identifying Practice Page 32 1
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Practice Page 32 1
  - User-Friendly Interface
4. Exploring eBook Recommendations from Practice Page 32 1
  - Personalized Recommendations
  - Practice Page 32 1 User Reviews and Ratings
  - Practice Page 32 1 and Bestseller Lists

5. Accessing Practice Page 32 1 Free and Paid eBooks
  - Practice Page 32 1 Public Domain eBooks
  - Practice Page 32 1 eBook Subscription Services
  - Practice Page 32 1 Budget-Friendly Options
6. Navigating Practice Page 32 1 eBook Formats
  - ePub, PDF, MOBI, and More
  - Practice Page 32 1 Compatibility with Devices
  - Practice Page 32 1 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Practice Page 32 1
  - Highlighting and Note-Taking Practice Page 32 1
  - Interactive Elements Practice Page 32 1
8. Staying Engaged with Practice Page 32 1
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Practice Page 32 1
9. Balancing eBooks and Physical Books Practice Page 32 1
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Practice Page 32 1
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Practice Page 32 1
  - Setting Reading Goals Practice Page 32 1
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Practice Page 32 1
  - Fact-Checking eBook Content of Practice Page 32 1
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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