

Physical Science Standards:

SPS1: Students will investigate our current understanding of the atom.

a. Examine the structure of the atom in terms of

- proton, electron, and neutron locations.
 - Proton has a charge of _____ and is located in the _____ of the atom.
 - Protons are also the elements' _____ number of the identity of the element, so protons are found by looking at the atomic number
 - Neutron has a charge of _____ and is located in the _____ of the atom.
 - Neutrons are found by subtracting the _____ Number from the _____ Mass or the BIG #- SMALL #.
 - Electron has a charge of _____ and is located in the _____ of the atom.
 - Electrons are equal to the number of _____ in a neutral atom.
 - The nucleus is made up of these 2 particles, _____ and _____.
 - The electrons are located in this region of the atom, known as the _____.
 - The electron cloud is divided into energy levels or energy shells, the first level will hold _____ electrons and the second level will hold _____ electrons.
 - The valence shell or valence level is the _____ energy shell of the atom.
- atomic mass and atomic number.
 - Atomic number is equal to the number of _____ in the atom and is also the atom's identity—it's the small number of the periodic table.
 - Atomic mass is equal to the average mass of all the isotopes of a particular element—it's the big number on the periodic table.
- atoms with different numbers of neutrons (isotopes).
 - Define Isotope _____.
 - Ex. Carbon-12 and Carbon-14
 - Carbon 12 has _____ P _____ N _____ E
 - Carbon 14 has _____ P _____ N _____ E
- explain the relationship of the proton number to the element's identity.
 - The _____ number is equal to the number of _____ in the atom which is how the element is identified.
 - Ex. Nitrogen has an atom number of _____, which means that it has _____ protons.

molybdenum	← element name
42	← atomic number number of protons (Z)
Mo	← atomic symbol
95.94	← atomic mass A (this is an average mass)



- Number of Protons = Atomic Number
- Number of Electrons = Number of Protons = Atomic Number
- Number of Neutrons = Mass Number - Atomic Number

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G Orfield



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Table of Contents Physical Science Eoct Study Guide Forsyth

1. Understanding the eBook Physical Science Eoct Study Guide Forsyth
 - The Rise of Digital Reading Physical Science Eoct Study Guide Forsyth
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Eoct Study Guide Forsyth
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Eoct Study Guide Forsyth
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Eoct Study Guide Forsyth
 - Personalized Recommendations
 - Physical Science Eoct Study Guide Forsyth User Reviews and Ratings
 - Physical Science Eoct Study Guide Forsyth and Bestseller Lists

5. Accessing Physical Science Eoct Study Guide Forsyth Free and Paid eBooks
 - Physical Science Eoct Study Guide Forsyth Public Domain eBooks
 - Physical Science Eoct Study Guide Forsyth eBook Subscription Services
 - Physical Science Eoct Study Guide Forsyth Budget-Friendly Options
6. Navigating Physical Science Eoct Study Guide Forsyth eBook Formats
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 - Physical Science Eoct Study Guide Forsyth Enhanced eBook Features
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 - Highlighting and Note-Taking Physical Science Eoct Study Guide Forsyth
 - Interactive Elements Physical Science Eoct Study Guide Forsyth
8. Staying Engaged with Physical Science Eoct Study Guide Forsyth
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Eoct Study Guide Forsyth
9. Balancing eBooks and Physical Books Physical Science Eoct Study Guide Forsyth
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Eoct Study Guide Forsyth
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Eoct Study Guide Forsyth
 - Setting Reading Goals Physical Science Eoct Study Guide Forsyth
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Eoct Study Guide Forsyth
 - Fact-Checking eBook Content of Physical Science Eoct Study Guide Forsyth
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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