

A chemical reaction is a transformation of a chemical substance in which the molecular formulae are written using chemical formulae instead of words.

Chemical reactions are written using chemical formulae instead of words.

A coefficient represents the number of atoms or molecules present in a chemical reaction. A coefficient is placed directly in front of a chemical formula.

A subscript is placed after the element symbol to indicate the number of atoms of that element in a chemical formula.

Why is it important to balance chemical equations?

Chemical reactions are balanced to show that the number of atoms of each element is the same on both sides of the equation. This is done by adding coefficients in front of the chemical formulae.

What does the total say is important to remember when balancing equations?

Remember, the total number of atoms of each element must be the same on both sides of the equation.

Week 8, Day 2

The number of protons in an atom cannot change (it is fixed) but a different number of neutrons can change.

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How does a positive ion form from a neutral atom? What is another name for a positive ion?

When an atom loses one or more electrons, it has more protons than it does electrons, so it is no longer neutral. This atom has a positive charge and is called a positive ion or cation.

How does a negative ion form from a neutral atom? What is another name for a negative ion?

When an atom gains one or more electrons, it has more electrons than it does protons, so it is no longer neutral. This atom has a negative charge and is called a negative ion or anion.

Atoms are formed by joining the first part of the element symbol and adding the suffix -ide. For example, sodium chloride, NaCl, is called a sodium ion.

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When positive and negative ions meet, the positive ions give up the electrons, becoming a neutral positively charged ion, and the negative ions become a neutral negatively charged ion.

The positively charged cation and the negatively charged anion are attracted to each other. This bond forms between the positively charged cation and the negatively charged anion.

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