

Chemical Equilibrium Lab

Introduction:

Some chemical systems reach the state of dynamic equilibrium involving interconverting reactants simultaneously forming as reactants or products and reactants. These reactions are called reversible reactions and the system can proceed in either direction. Once the equilibrium is reached the rates of the forward and reverse reactions are equal. Equilibrium is the point at which the concentrations of reactants and products are constant. The importance of equilibrium is that equilibrium is dynamic and the equilibrium is not static. Reactions that reach equilibrium can be shifted. Le Chatelier's Principle states that if an equilibrium system is subjected to a change in one of its conditions such as temperature or concentration or pressure or volume or the amount of one of the reactants or products or the equilibrium constant it will shift in a direction that will counteract the change and bring the system back to equilibrium. When more products are added an equilibrium reaction will shift to the left to consume the added products and bring the system back to equilibrium. When more reactants are added an equilibrium reaction will shift to the right to consume the added reactants and bring the system back to equilibrium. In this lab, you will be observing the changes in equilibrium when the amount of reactants, and temperature are changed. Follow Le Chatelier's Principle in the following reactions. Shifted equilibrium is indicated by the color change or solution becoming by the color indicators in the equations below. For example, if the color changes to red, the reaction shifts to the right or shifts to the right or shifts to the right.



Adding FeCl_3 increases $\text{Fe}^{3+}(\text{aq})$ and will shift the $\text{Fe}^{3+}(\text{aq})$ to the right, making the solution more red. Adding KSCN will shift the $\text{SCN}^{-}(\text{aq})$ to the right, making the solution more red.

Procedure (This Group):

1. Prepare 100 mL of 0.1 M FeCl_3 solution.
2. Prepare 100 mL of 0.1 M KSCN solution.
3. Prepare 100 mL of 0.1 M FeCl_3 solution.
4. Prepare 100 mL of 0.1 M KSCN solution.
5. Prepare 100 mL of 0.1 M FeCl_3 solution.
6. Prepare 100 mL of 0.1 M KSCN solution.
7. Prepare 100 mL of 0.1 M FeCl_3 solution.
8. Prepare 100 mL of 0.1 M KSCN solution.
9. Prepare 100 mL of 0.1 M FeCl_3 solution.
10. Prepare 100 mL of 0.1 M KSCN solution.
11. Prepare 100 mL of 0.1 M FeCl_3 solution.
12. Prepare 100 mL of 0.1 M KSCN solution.
13. Prepare 100 mL of 0.1 M FeCl_3 solution.
14. Prepare 100 mL of 0.1 M KSCN solution.
15. Prepare 100 mL of 0.1 M FeCl_3 solution.
16. Prepare 100 mL of 0.1 M KSCN solution.
17. Prepare 100 mL of 0.1 M FeCl_3 solution.
18. Prepare 100 mL of 0.1 M KSCN solution.
19. Prepare 100 mL of 0.1 M FeCl_3 solution.
20. Prepare 100 mL of 0.1 M KSCN solution.

Observing Chemical Equilibrium Lab 46

Sabine Zange



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Watching Ultrafast Molecular Motions with 2D IR Chemical Exchange Spectroscopy Michael D. Fayer, 2011 This unique volume presents a comprehensive but accessible introduction to the field of ultrafast two dimension infrared 2D IR vibrational echo spectroscopy based on the pioneering work of Professor Michael D Fayer Department of Chemistry Stanford University USA It contains in one place a qualitative introduction to the field of 2D IR spectroscopy and a comprehensive set of scientific papers that underlie the qualitative discussion The introductory material contains several detailed illustrations and is based on the Centenary Lecture at the Indian Institute of Science given by Professor Fayer July 16 2008 as part of the celebration of the 100th anniversary of the founding of IIS in Bangalore India The second part of the volume contains reprints of Fayer's relevant papers The compilation will be very useful because it presents the historical background motivation methodology and experimental results at a level that is accessible to the non expert The reprints of the scientific papers from review articles to detailed theoretical papers provide rigorous supporting material so that the reader can delve as deeply as desired into the subject **Observations of Surface**

to Atmosphere Interactions in the Tropics Michael Garstang, David Roy Fitzjarrald, 1999 Called the firebox of the atmosphere the tropics absorb more energy from the Sun than they lose through longwave emissions this excess energy activates processes in the temperate and polar regions of the Earth This book documents the historical evolution of concepts which describe the complex interactions of scales of motion which connect the surface mixed and cloud layers to the deeper

atmosphere of the tropics Thermodynamic and kinematic consequences of these transfers of energy are extended to the geochemical and living worlds **Experiment Station Record** United States. Office of Experiment Stations,1913

Observation, Prediction and Simulation of Phase Transitions in Complex Fluids Marc Baus,L.F Rull,Jean-Paul Ryckaert,2012-12-06 Observation Prediction and Simulation of Phase Transitions in Complex Fluids presents an overview of the phase transitions that occur in a variety of soft matter systems colloidal suspensions of spherical or rod like particles and their mixtures directed polymers and polymer blends colloid polymer mixtures and liquid forming mesogens This modern and fascinating branch of condensed matter physics is presented from three complementary viewpoints The first section written by experimentalists emphasises the observation of basic phenomena by light scattering for example The second section written by theoreticians focuses on the necessary theoretical tools density functional theory path integrals free energy expansions The third section is devoted to the results of modern simulation techniques Gibbs ensemble free energy calculations configurational bias Monte Carlo The interplay between the disciplines is clearly illustrated For all those interested in modern research in equilibrium statistical mechanics Experiment Station Record U.S. Office of Experiment Stations,United States. Agricultural Research Service,United States. Office of Experiment Stations,1913 *The Iron(III) Thiocyanate Reaction* Kevin C. de Berg,2019-11-13 This Brief presents an historical investigation into the reaction between ferric ions and thiocyanate ions which has been viewed in different ways throughout the last two centuries Historically the reaction was used in chemical analysis and to highlight the nature of chemical reactions the laws of chemistry models and theories of chemistry chemical nomenclature mathematics and data analysis and instrumentation which are important ingredients of what one might call the nature of chemistry Using the history of the iron III thiocyanate reaction as a basis the book s main objective is to explore how chemistry develops its own knowledge base how it assesses the reliability of that base and how some important tools of the trade have been brought to bear on a chemical reaction to achieve understanding a worthwhile goal of any historical investigation **Essential Practical NMR for Organic Chemistry** S. A. Richards,J. C. Hollerton,2023-04-03 Essential Practical NMR for Organic Chemistry A hands on resource advocating an ordered approach to gathering and interpreting NMR data The second edition of Essential Practical NMR for Organic Chemistry delivers a pragmatic and accessible text demonstrating an ordered approach to gathering and interpreting NMR data In this informal guide you ll learn to make sense of the high density of NMR information through the authors problem solving strategies and interpretations The book also discusses critical aspects of NMR theory as well as data acquisition and processing strategy It explains the use of NMR spectroscopy for dealing with problems of small organic molecule structural elucidation and includes a brand new chapter on Nitrogen 15 NMR Readers will also find Strategies for preparing a sample spectrum acquisition processing and interpreting your spectrum Fulsome discussions of Carbon 13 NMR spectroscopy Practical treatments of quantification safety procedures and relevant software An ideal handbook for anyone involved in using NMR to

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Scientific and Technical Aerospace Reports ,1995 Lists citations with abstracts for aerospace related reports obtained from world wide sources and announces documents that have recently been entered into the NASA Scientific and Technical Information Database Acid Precipitation ,1994 *Publications of the National Institute of Standards and Technology ... Catalog* National Institute of Standards and Technology (U.S.),1991

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