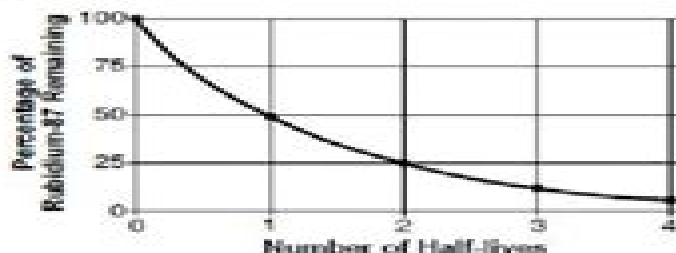


Name: \_\_\_\_\_

Answer all questions on the answer document. This is a TAKE HOME TEST due on Wednesday, June 7.

1. The graph below shows the radioactive decay of rubidium-87.



What percentage of rubidium-87 atoms will be left after four half-lives?

- A. 25.0%  
 B. 12.5%  
 C. 6.25%  
 D. 3.125%
2. Given the equation representing a reaction at equilibrium:
- $$\text{H}_2\text{S}(\text{aq}) + \text{CH}_3\text{NH}_2(\text{aq}) \rightleftharpoons \text{HS}^-(\text{aq}) + \text{CH}_3\text{NH}_3^+(\text{aq})$$
- According to one acid-base theory, the forward reaction is classified as an acid-base reaction because
- A.  $\text{H}_2\text{S}$  is a  $\text{H}^+$  donor and  $\text{CH}_3\text{NH}_2$  is a  $\text{H}^+$  acceptor  
 B.  $\text{CH}_3\text{NH}_2$  is a  $\text{H}^+$  donor and  $\text{H}_2\text{S}$  is a  $\text{H}^+$  acceptor  
 C.  $\text{HS}^-$  and  $\text{CH}_3\text{NH}_3^+$  are both  $\text{H}^+$  donors  
 D.  $\text{CH}_3\text{NH}_3^+$  and  $\text{HS}^-$  are both  $\text{H}^+$  acceptors
3. Which ionic equation represents a spontaneous reaction that can occur in a voltaic cell?
- A.  $\text{Cu}(\text{s}) + \text{Zn}(\text{s}) \rightarrow \text{Cu}^{2+}(\text{aq}) + \text{Zn}^{2+}(\text{aq})$   
 B.  $\text{Cu}(\text{s}) + \text{Zn}^{2+}(\text{aq}) \rightarrow \text{Cu}^{2+}(\text{aq}) + \text{Zn}(\text{s})$   
 C.  $\text{Cu}^{2+}(\text{aq}) + \text{Zn}(\text{s}) \rightarrow \text{Cu}(\text{s}) + \text{Zn}^{2+}(\text{aq})$   
 D.  $\text{Cu}^{2+}(\text{aq}) + \text{Zn}^{2+}(\text{aq}) \rightarrow \text{Cu}(\text{s}) + \text{Zn}(\text{s})$

4. Given the equation representing a reaction:
- $$\text{Sn}^{4+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Sn}^{2+}(\text{aq})$$

Which term best describes this reaction?

- A. ionization  
 B. neutralization  
 C. oxidation  
 D. Reduction

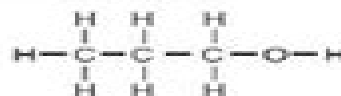
5. The table below shows the atomic mass and natural abundance of the two naturally occurring isotopes of lithium.

| Naturally Occurring Isotopes of Lithium |                 |                       |
|-----------------------------------------|-----------------|-----------------------|
| Isotope                                 | Atomic Mass (u) | Natural Abundance (%) |
| Li-6                                    | 6.015           | 7.6                   |
| Li-7                                    | 7.016           | 92.4                  |

Which numerical setup can be used to determine the atomic mass of naturally occurring lithium?

- A.  $(7.6)(6.015 \text{ u}) + (92.4)(7.016 \text{ u})$   
 B.  $(0.076)(6.015 \text{ u}) + (0.924)(7.016 \text{ u})$   

$$\frac{(7.6)(6.015 \text{ u}) + (92.4)(7.016 \text{ u})}{2}$$
  
 C. 
$$\frac{(0.076)(6.015 \text{ u}) + (0.924)(7.016 \text{ u})}{2}$$
  
 D. 
$$\frac{(7.6)(6.015 \text{ u}) + (92.4)(7.016 \text{ u})}{2}$$
6. Given the formulas representing two compounds at standard pressure:



1-propanol



methyl ethyl ether

The compounds can be differentiated by their

- A. boiling points  
 B. gram-formula masses  
 C. numbers of hydrogen atoms  
 D. percent compositions by mass of carbon
7. Given the formula representing a compound:



What is the IUPAC name of this compound?

- A. 2-chloroheptane  
 B. 6-chloroheptane  
 C. 2,2-dichloroheptane  
 D. 6,6-dichloroheptane

# Regents Chemistry The Periodic Law Worksheet

**Effiong Eyo**



## **Regents Chemistry The Periodic Law Worksheet:**

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