

BALANCING REDOX REACTIONS

STEP BY STEP IN TWO WAYS...



BOTH START THE SAME WAY

1. ASSIGN OXIDATION NUMBERS TO EACH ATOM



NOTE: OXIDATION NUMBERS SUM UP TO ZERO EVERYWHERE

2. IDENTIFY THE OXIDATION AND REDUCTION HALF-REACTIONS

3. BALANCE EACH HALF-REACTION SEPARATELY

4. REVERSE YOUR EQUATION AND THE HALF-REACTIONS BALANCING ALL ATOMS INVOLVED IN IT

NOTE: REVERSE IT! + BECOMES - AND VICE-VERSA



HALF-REACTION METHOD (in acidic conditions)

5. BALANCE OXIDATION



6. BALANCE REDUCTION



7. BALANCE CHARGES WITH ELECTRONS

NOTE: BALANCE CHARGES FIRST BY ADDING H⁺ AND H₂O THEN THE ELECTRONS



8. MAKE SURE THE NUMBER OF ELECTRONS IS THE SAME

NOTE: YOU CAN BALANCE THIS LAST STEP FIRST, THEN ADD H⁺ AND H₂O



9. PUT BOTH REACTIONS BACK TOGETHER



10. CHECK FOR ANYTHING THAT IS NOT A PRODUCT AND CANCELLED

NOTE: OTHER REACTIONAL PARTS MUST STAY AND CANCEL OUT



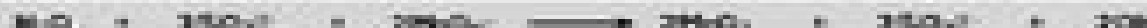
BUT WHAT ABOUT BASIC CONDITIONS?

WHEN WE SAY "ACIDIC" OR "BASIC" WE ARE USING TWO APPROXIMATE DEFINITIONS:

ACIDIC CONDITIONS HAVE H⁺ AND BASIC CONDITIONS HAVE OH⁻

SO WHEN WE FIND OUR REACTION FROM ABOVE, WITH A REACTION PERFORMED IN BASIC CONDITIONS:

8. ADD THE NUMBER OF OH⁻ TO THE NUMBER OF H⁺ IN EACH SIDE
NOTE: YOU CAN DO THIS LAST STEP FIRST, THEN ADD H⁺ AND H₂O



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Minjie Lin



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