

Assessment 17 Worksheet: Redox Reactions & Oxidation Numbers

Complete the following questions, and upload the file to the dropbox.

- Why can an oxidation reaction not happen without a reduction reaction? Your response should clearly communicate your understanding of the nature of redox reactions. (1 point)
An oxidation reaction can not happen without a reduction reaction because for an oxidation reaction there must also be a reduction happening at the same time. This is because oxidation happens because of the "compound" that causes the oxidation, this "compound" must be reduced otherwise oxidation will not occur.
- Determine the oxidation number of the specified element in each of the following. (6 points)
 - Nitrogen in NF_3 +3
 - Sulfur in S_8 0
 - Chromium in CrO_4^{2-} +5
 - Phosphorus in P_2O_5 +5
 - Carbon in $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ 0
 - Carbon in CHCl_3 +2
- Determine the oxidation number for each element in the following substances. (6 points)
 - H_2SO_3 hydrogen: +1 sulfur: +4 oxygen: -2
 - OH^- oxygen: -2 hydrogen: +1
 - HPO_4^{2-} hydrogen: +1 phosphorus: +5 oxygen: -2
- Determine whether each reaction is a redox reaction by finding the oxidation numbers. Show the oxidation numbers to justify your answer. (6 points)
 - $$\overset{+1 \quad -1}{\text{H}_2\text{O}_2} + 2 \overset{+2 \quad -2 \quad +1}{\text{Fe}(\text{OH})_2} \rightarrow 2 \overset{+3 \quad -2 \quad +1}{\text{Fe}(\text{OH})_3}$$

$$\text{H}_2\text{O}_2 + 2 \text{Fe}(\text{OH})_2 \rightarrow 2 \text{Fe}(\text{OH})_3$$

| < oxidation > |
 - $$\overset{+3 \quad -1}{\text{PCl}_3} + 3 \overset{+1 \quad -2}{\text{H}_2\text{O}} \rightarrow \overset{+1 \quad +3 \quad -2}{\text{H}_3\text{PO}_3} + 3 \overset{+1 \quad +1}{\text{HCl}}$$

$$\text{PCl}_3 + 3 \text{H}_2\text{O} \rightarrow \text{H}_3\text{PO}_3 + 3 \text{HCl}$$

Redox or Not? Redox

Redox or Not? Not a Redox

Redox Reactions Chapter Assessment Answers

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