

Chem 12 Electrochemistry : Quiz 2 - 120

Part 1

1) A particular reaction has a predicted E° of -2.25 V. This indicates that the reaction would :

- a) occur rapidly b) produce energy c) go to completion
d) be non-spontaneous

2) Corrosion of a metal means that the metal :

- a) acts as a cathode b) undergoes oxidation
c) undergoes reduction d) acts as an oxidizing agent

3) Zinc metal is prepared by electrolysis as follows:



What is the minimum voltage necessary for the electrolysis reaction to proceed ?

- a) 2.12 b) 1.58 c) 1.96 d) 3.48

4) Which metal would provide the best cathodic protection for iron pipes ?

- a) Ag b) Sn c) Mg d) Cu

5) Which half-reaction best describes what is taking place at the anode during the electrolysis of a 1.0 M NaCl solution ?

- a) $\text{Na}^{+} + \text{e}^{-} \rightarrow \text{Na}$ b) $2\text{Cl}^{-} \rightarrow \text{Cl}_2(\text{g}) + 2\text{e}^{-}$
c) $2\text{H}^{+}(10^{-7}\text{M}) + 2\text{e}^{-} \rightarrow \text{H}_2(\text{g})$ d) $\text{H}_2\text{O} \rightarrow 1/2 \text{O}_2(\text{g}) + 2\text{H}^{+}(10^{-7}\text{M}) + 2\text{e}^{-}$

6) Which one of the following converts chemical energy to electrical energy ?

- a) Industrial production of aluminum b) Nickel plating
c) Discharging a "dry" cell d) Production of Na(s)

7) Which is the unit for the energy per coulomb of electrons ?

Redox And Electrochemistry Quiz Answers

H Kauffman



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