

Elements of the Sea

Ionic equations

- a. If 2 solutions react to form a solid, a precipitation reaction occurs. Eg. $\text{Ag}^+(\text{aq}) + \text{NO}_3^-(\text{aq}) + \text{Na}^+(\text{aq}) + \text{Br}^-(\text{aq})$ goes to $\text{AgBr}(\text{s}) + \text{Na}^+(\text{aq}) + \text{NO}_3^-(\text{aq})$

Then cancel spectators

$\text{Ag}^+(\text{aq}) + \text{Br}^-(\text{aq})$ goes to $\text{AgBr}(\text{s})$

- a. Always include state symbols
- a. Only split into ions if **aqueous and ionic**

Ionic solids

- a. Cations and anions are held together by strong electrostatic attractions in a **giant ionic lattice**.
- a. Each sodium ion is surrounded by 6 chloride ions **6:6 coordination**. It is a **simple cubic** structure.
- a. Ionic solids are hard and have a high melting point due to strong electrostatic attractions between ions.
- a. If ionic crystals contain water, the water molecules fit into the lattice the same way as regular ions. This is called the **water of crystallisation**. These crystals are said to be hydrated.
- a. If these hydrated crystals are heated, the water molecules are driven off and you are left with an **anhydrous** solid.



Ionic substances in solution

- a. As ionic substances dissolve in water, the ions become surrounded by water molecules and they spread throughout the solution. The hydrated ions are **randomly** arranged and behave **independently**.

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