

BALANCING REDOX REACTIONS

Oxidation Reduction



- During **oxidation**, oxidation number increases.
- During **reduction**, oxidation number decreases.

The reduced reactant is the oxidizing agent.

- H^+ gains electrons and changes to hydrogen gas by taking electrons from it. (its o.n. decreases from +1 to 0)

The oxidized reactant is the reducing agent.

- Zn loses electrons and changes to zinc ions. (its o.n. increases from 0 to +2)

High School Chemistry
Lesson Slides



Redox Reactions High School Chemistry Activity

BM King

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular gradient shape partially visible behind it.

Redox Reactions High School Chemistry Activity:

Reviewing **Redox Reactions High School Chemistry Activity**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Redox Reactions High School Chemistry Activity**," an enthralling opus penned by a very acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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