

CHAPTER

08

REDOX REACTION

- + The chemical reaction which contain oxidation as well as reduction half reaction is called redox reaction.
- + An atom or ion or molecule which undergoes reduction (gain of electrons) is called oxidant (while an atom or molecule which undergoes oxidation) (loss of electrons) is called reductant.
- + Oxygen shows positive oxidation states (+1 and +2) only in OF_2 and O_2F_2 .
- + Maximum oxidation state of an element never more than its group number.
- + a) Oxidation state of Mn in $KMnO_4$ is +7.
b) Oxidation state of Cr in $K_2Cr_2O_7$ is +6.
c) Oxidation state of Cr in CrO_3 is +6.
- + a) Oxidation state of 'N' in N_3H is -1/3.
b) Oxidation state of 'N' in NH_4NO_3 (NH_4^+ and NO_3^-) are -3 and +5.
- + a) Oxidation state of 'Fe' in $FeSO_4$ is +2.
b) Oxidation state of 'Fe' in Fe_3O_4 is +8/3.
- + In $xMnO_4^- + yH_2O + zI^- \rightarrow 2MnO_2 + 8OH^- + 3I_2$
 $x = 2, y = 4, z = 6$
- + In reaction of $P_4 + xH_2O + yOH^- \rightarrow PH_3 + zH_2PO_2^-$
 $x = 3, y = 3, z = 3$ (i.e. $x = y = z$)
- + $CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$ is an example for chemical combustion.
- + In all combustion reactions oxygen undergoes reduction (0 to -2).
- + In decomposition reaction chemical compound split into two or more simpler substances.
 $[2H_2O \rightarrow 2H_2 + O_2]$
- + In decomposition of $MgCO_3$, no changes in oxidation numbers. So it is not a redox reaction
 $[MgCO_3 \rightarrow MgO + CO_2]$
- + $Zn_{(s)} + CuSO_{4(aq)} \rightarrow ZnSO_{4(aq)} + Cu_{(s)}$ is a displacement reaction.
- + In disproportionation reaction the same element undergoes both oxidation and reduction simultaneously.
- + When Br_2 and I_2 dissolve in CCl_4 solvent to give reddish brown and purple colours respectively. So, they are detected by extracting into CCl_4 . This is also called "Layer Test".
- + Comproportion is a reverse of disproportionation reaction

Redox Reactions Chapter Study

**J.K. Gilbert, Onno de Jong, Rosária
Justi, David F. Treagust, Jan H. van Driel**



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Gilbert, Onno de Jong, Rosária Justi, David F. Treagust, Jan H. van Driel, 2006-03-11 Chemical education is essential to everybody because it deals with ideas that play major roles in personal social and economic decisions This book is based on

three principles that all aspects of chemical education should be associated with research that the development of opportunities for chemical education should be both a continuous process and be linked to research and that the professional development of all those associated with chemical education should make extensive and diverse use of that research It is intended for pre service and practising chemistry teachers and lecturers chemistry teacher educators chemical education researchers the designers and managers of formal chemical curricula informal chemical educators authors of textbooks and curriculum support materials practising chemists and chemical technologists It addresses the relation between chemistry and chemical education curricula for chemical education teaching and learning about chemical compounds and chemical change the development of teachers the development of chemical education as a field of enquiry This is mainly done in respect of the full range of formal education contexts schools universities vocational colleges but also in respect of informal education contexts books science centres and museums

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