Redox: IB Review Questions

Consider the standard electrode potentials of the following reactions:

$$\operatorname{Sn}^{4}{}^{\circ}(\operatorname{aq}) + 2e^{-} \rightarrow \operatorname{Sn}^{2}{}^{\circ}(\operatorname{aq})$$
 +0.15 V

$$Fe^{3+}(aq) + e^{-} \rightarrow Fe^{2+}(aq)$$
 +0.77 V

What is the value of the cell potential (in volts) for the spontaneous reaction?

A. +1.69

- $B_{\rm c} + 1.39$
- C. +0.92
- D. +0.62
- Which factors affect the amount of metal formed during electrolysis? 2.
 - E. Charge on the metal ion
 - Current
 - HD. Time
 - Acres Land II only

ER. I amd III conly

C. III and III only

- D. I, III and III
- Consider the following reactions.

$$Cu^{2+}(aq) + 2e^- \rightleftharpoons Cu(s)$$
 $E^+ = +0.34 \text{ V}$

$$Mg^{2*}(aq) + 2e^{-} \rightleftharpoons Mg(s) \quad E^{*} = -2.36 \text{ V}$$

$$Zn^{2^{*}}(sq) + 2e^{-} \rightleftharpoons Zn(s)$$
 $E^{*} = -0.76 \text{ V}$

Which statement is correct?

- Cu2 (aq) will oxidize both Mg(s) and Zn(s).
- Zn(s) will reduce both Cu2 (aq) and Mg2 (aq).
- Mg2 (aq) will oxidize both Cu(s) and Zn(s). C.
- Cu(s) will reduce both Mg2" (aq) and Zn2" (aq).
- Consider the standard electrode potentials of the following reactions.

$$Cr^{3+}(aq) + 3e^- \rightarrow Cr(s)$$
 -0.75 V

$$Cd^{2+}(aq) + 2e^- \rightarrow Cd(s)$$
 -0.40 V

What is the value of the cell potential (in V) for the following reaction?

$$2Cr(s) + 3Cd^{2s}(aq) \rightarrow 2Cr^{1s}(aq) + 3Cd(s)$$

- A. -0.35 B. -1.15 C. +0.30
- 80% +00.335

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Joacim Rocklöv

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Table of Contents Redox Ib Review Questions And Answers

- 1. Understanding the eBook Redox Ib Review Questions And Answers
 - The Rise of Digital Reading Redox Ib Review Questions And Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Redox Ib Review Questions And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Redox Ib Review Questions And Answers
 - User-Friendly Interface

- 4. Exploring eBook Recommendations from Redox Ib Review Questions And Answers
 - Personalized Recommendations
 - Redox Ib Review Questions And Answers User Reviews and Ratings
 - Redox Ib Review Questions And Answers and Bestseller Lists
- 5. Accessing Redox Ib Review Questions And Answers Free and Paid eBooks
 - Redox Ib Review Ouestions And Answers Public Domain eBooks
 - Redox Ib Review Questions And Answers eBook Subscription Services
 - Redox Ib Review Questions And Answers Budget-Friendly Options
- 6. Navigating Redox Ib Review Questions And Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - Redox Ib Review Questions And Answers Compatibility with Devices
 - Redox Ib Review Questions And Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Redox Ib Review Questions And Answers
 - Highlighting and Note-Taking Redox Ib Review Questions And Answers
 - Interactive Elements Redox Ib Review Questions And Answers
- 8. Staying Engaged with Redox Ib Review Questions And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Redox Ib Review Questions And Answers
- 9. Balancing eBooks and Physical Books Redox Ib Review Questions And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Redox Ib Review Questions And Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Redox Ib Review Questions And Answers
 - Setting Reading Goals Redox Ib Review Questions And Answers
 - Carving Out Dedicated Reading Time

- 12. Sourcing Reliable Information of Redox Ib Review Questions And Answers
 - Fact-Checking eBook Content of Redox Ib Review Questions And Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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