

**Balancing a Redox Equation for an Acidic Solution
Using the Oxidation Number Method**
(Student textbook page 615)

35. Use the oxidation number method to balance the following equation for the combustion of carbon disulfide:



What Is Required?

You must balance an equation for the combustion of carbon disulfide, using the oxidation number method.

What Is Given?

You are given the reactants, carbon disulfide and oxygen, and the products, carbon dioxide and sulfur dioxide.

You are given rules for assigning oxidation numbers in **Table 9.3** on page 604 of the student textbook.

Plan Your Strategy	Act on Your Strategy
Step 1 Assign an oxidation number to each atom in the equation and determine whether it is a redox reaction.	$\overset{+4 \ -2}{\text{CS}_2}(\text{g}) + \overset{0}{\text{O}_2}(\text{g}) \rightarrow \overset{+4 \ -2}{\text{CO}_2}(\text{g}) + \overset{+4 \ -2}{\text{SO}_2}(\text{g})$ The oxidation numbers of sulfur and oxygen change during the reaction. It is a redox reaction.
Step 2 Identify the atom or atoms that undergo an increase in oxidation number and the atom or atoms that undergo a decrease in oxidation number.	Each of two sulfur atoms lost 6 electrons and each of two oxygen atoms gained 2 electrons. So the O_2 lost 12 electrons and the CS_2 gained 4 electrons. $ \begin{array}{c} \text{1}(-12\text{e}^-) = -12\text{e}^- \\ \downarrow \\ \overset{+4 \ -2}{\text{CS}_2}(\text{g}) + \overset{0}{\text{O}_2}(\text{g}) \rightarrow \overset{+4 \ -2}{\text{CO}_2}(\text{g}) + \overset{+4 \ -2}{\text{SO}_2}(\text{g}) \\ \uparrow \qquad \qquad \qquad \uparrow \qquad \qquad \uparrow \\ \text{3}(+4\text{e}^-) = +12\text{e}^- \end{array} $

Redox Reaction Solutions Manual

Joseph Laurino



Redox Reaction Solutions Manual:

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Principles of Physical Chemistry Third Edition This book provides a unique approach to introduce undergraduate students to the concepts and methods of physical chemistry which are the foundational principles of Chemistry The book introduces the student to the principles underlying the essential sub fields of quantum mechanics atomic and molecular structure atomic and molecular spectroscopy statistical thermodynamics classical thermodynamics solutions and equilibria electrochemistry kinetics and reaction dynamics macromolecules and organized molecular assemblies Importantly the book develops and applies these principles to supramolecular assemblies and supramolecular machines with many examples from biology and nanoscience In this way the book helps the student to see the frontier of modern physical chemistry developments The book begins with a discussion of wave particle duality and proceeds systematically to more complex chemical systems in order to relate the story of physical chemistry in an intellectually coherent manner The topics are organized to correspond with those typically given in each of a two course semester sequence The first 13 chapters present quantum mechanics and spectroscopy to describe and predict the structure of matter atoms molecules and solids Chapters 14 to 29 present statistical thermodynamics and kinetics and applies their principles to understanding equilibria chemical transformations macromolecular properties and supramolecular machines Each chapter of the book begins with a simplified view of a topic and evolves to more rigorous description in order to provide the student and instructor flexibility to choose the level of rigor and detail that suits them best The textbook treats important new directions in physical chemistry research including chapters on macromolecules principles of interfaces and films for organizing matter and supramolecular machines as well as including discussions of modern nanoscience spectroscopy and reaction dynamics throughout the text *Student Solutions Manual for Exploring Chemical Analysis* Daniel C. Harris, 2008-04-18 Exploring Chemical Analysis teaches students how to understand analytical results and how to use quantitative manipulations preparing them for the problems they will encounter

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Solutions Manual to Accompany Physical Chemistry for the Life Sciences C. A. Trapp, M. P. Cady, 2011 This solutions manual contains fully worked solutions to all end of chapter discussion questions and exercises featured in Physical Chemistry for the Life Sciences

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sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter **Solutions**

Manual for Chemistry: Molecules Matter and Change, Fourth Edition Julie Henderleiter, Charles Trapp, Peter Atkins, Loretta Jones, 1999-08-13 This student companion is a supplement to Chemistry Molecules Matter and Change 4th edition with CD ROM It features guided reading strategies collaborative learning sheets and strategies for using CD ROM tools **US Solutions Manual to Accompany Elements of Physical Chemistry 7e** David Smith, 2017-09-28 The Solutions Manual to Accompany Elements of Physical Chemistry 7th edition contains full worked solutions to all end of chapter discussion questions and exercises featured in the book The manual provides helpful comments and friendly advice to aid understanding It is also a valuable resource for any lecturer who wishes to use the extensive selection of exercises featured in the text to support either formative or summative assessment and wants labour saving ready access to the full solutions to these questions Study Guide and Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry, Third Edition John McMurry, Susan McMurry, 1999 Provides worked out solutions to text problems along with chapter by chapter outlines and a variety of self tests at the end of each chapter *Student Solutions Manual for Chang's Chemistry* Philip C. Keller, Raymond Chang, Jill L. Keller, 1988 Laboratory Manual in Engineering Chemistry : For the Students of JNTU Hyderabad Dr. B. Rama Devi & Dr. P. Aparna, The book is written to gain the basic knowledge on the principles of chemistry required for practical applications in engineering concepts This book consists organic and general chemistry experiments for chemical engineering for 1st and 2nd semester students The book also explains the precautions and safety rules for avoiding the accidents in chemistry laboratory It covers Estimation of Ferrous iron by Dichrometry and Permanganometry Method Estimation of Acetic Acid by Conductometric Titrations Estimation of the Amount of Fe²⁺ by Potentiometry Determination of an Acid Concentration using pH Meter Preparation of Nylon 6 and Bakelite Phenol Formaldehyde Resin Estimation of Acid Value of Given Lubricant Oil Determination of Rate of Corrosion of Mild Steel Preparation of Benzanilide from Benzophenone via the Oxime by Beckmann Rearrangement etc

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