

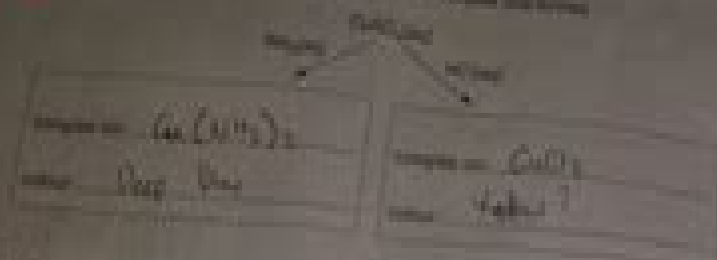
Question 1

Answer all the questions

The reaction is about the synthesis of copper complexes and ligands.

(a) The reaction shows the synthesis of a copper complex.

In the boxes, write the formula and name of the ligands and the complex.



(b) Cu^{2+} ions form a complex with two ethylenediamine ligands and two water molecules. The ethylenediamine ligand is a bidentate ligand.

The skeletal formula of the ethylenediamine ligand is shown in Fig. 1.1 below.



Fig. 1.1

(c) What is meant by the term bidentate ligand?

A ligand with two donor atoms forming two coordinate bonds with central metal ion.

(d) The complex is shown in Fig. 1.2 below.

The structure of the ethylenediamine ligand is shown below. The charge has been added.



Complex A

Complete the table by giving the other two geometries of A. You do not need to include any charges.

Indicate with ticks whether the stereoisomers are cis, trans, optical or a combination of these types.

Structure														
Type	<table><tr><td>cis</td><td></td></tr><tr><td>trans</td><td>✓</td></tr><tr><td>optical</td><td></td></tr></table>	cis		trans	✓	optical		<table><tr><td>cis</td><td>✓</td></tr><tr><td>trans</td><td></td></tr><tr><td>optical</td><td>✓</td></tr></table>	cis	✓	trans		optical	✓
cis														
trans	✓													
optical														
cis	✓													
trans														
optical	✓													

(e) What is the empirical formula, including the charge, of the complex in A?

$[Cu(NH_3)_2(en)_2]^{2+}$

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R Barnett



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