

Question number	Answer	Marks	Guidance
1	$\text{Ca} : \text{N} : \text{O} = 30.35/40.1 : 21.20/14.0 : 48.45/16.0$ $= 0.7569 : 1.5286 : 3.028$ Formula = CaN_2O_4	B1 B1	
2 (a)	$2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$	B1	
2 (b)	$4.50 \times 10^{-3} \text{ mol}$	B1	
2 (c)	$n(\text{KClO}_3) = 3.0 \times 10^{-3} \text{ mol}$ $M(\text{KClO}_3) = 122.6 \text{ g mol}^{-1}$ mass of $\text{KClO}_3 = 0.3678 \text{ g}$	B1 B1 B1	
3	$M(\text{ZnSO}_4) = 161.5 \text{ g mol}^{-1}$ $n(\text{ZnSO}_4) = 6.57 \times 10^{-3} \text{ mol}$ $n(\text{H}_2\text{O}) = (1.893 - 1.061)/18 = 4.62 \times 10^{-2} \text{ mol}$ $x = n(\text{H}_2\text{O})/n(\text{ZnSO}_4) = 7$ Formula = $\text{ZnSO}_4 \cdot 7\text{H}_2\text{O}$	B1 B1 B1 B1	
4	$M(\text{Na}_2\text{CO}_3) = 106.0 \text{ g mol}^{-1}$ Actual $n(\text{Na}_2\text{CO}_3) = 0.0234 \text{ mol}$ Theoretical $n(\text{Na}_2\text{CO}_3) = 0.0234 \times 100/65$ $= 0.0360 \text{ mol}$ Theoretical $n(\text{NaHCO}_3) = 0.0720 \text{ mol}$ Mass of $\text{NaHCO}_3 = 0.0720 \times 84.0 = 6.05 \text{ g}$	B1 B1 B1 B1	
5 (a)	$\text{Fe}_2\text{O}_3(\text{s}) + 3\text{CO}(\text{g}) \rightarrow 2\text{Fe}(\text{s}) + 3\text{CO}_2(\text{g})$	B1	
5 (b)	$M(\text{Fe}_2\text{O}_3) = 159.6 \text{ g mol}^{-1}$ $n(\text{Fe}_2\text{O}_3) = (10\,000 \times 10^6)/159.6 = 6.266 \times 10^7 \text{ mol}$ $n(\text{Fe}) = 1.253 \times 10^8 \text{ mol} = 6.992 \times 10^5 \text{ g}$ (6992 tonnes)	B1 B1 B1	
6 (a)	$\text{Na} : \text{N} : \text{O} = 27.1/23.0 : 16.5/14.0 : 56.4/16.0$ $= 1.178 : 1.179 : 3.525$ Formula = NaNO_3	B1 B1	
6 (b)	$2\text{NaNO}_3(\text{s}) \rightarrow 2\text{NaNO}_2(\text{s}) + \text{O}_2(\text{g})$	B1	

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DJ Losen



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Table of Contents Ocr Chemistry Exam Question Answers

1. Understanding the eBook Ocr Chemistry Exam Question Answers
 - The Rise of Digital Reading Ocr Chemistry Exam Question Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Chemistry Exam Question Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr Chemistry Exam Question Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Chemistry Exam Question Answers
 - Personalized Recommendations
 - Ocr Chemistry Exam Question Answers User Reviews and Ratings
 - Ocr Chemistry Exam Question Answers and Bestseller Lists
5. Accessing Ocr Chemistry Exam Question Answers Free and Paid eBooks
 - Ocr Chemistry Exam Question Answers Public Domain eBooks

- Ocr Chemistry Exam Question Answers eBook Subscription Services
- Ocr Chemistry Exam Question Answers Budget-Friendly Options
- 6. Navigating Ocr Chemistry Exam Question Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Ocr Chemistry Exam Question Answers Compatibility with Devices
 - Ocr Chemistry Exam Question Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ocr Chemistry Exam Question Answers
 - Highlighting and Note-Taking Ocr Chemistry Exam Question Answers
 - Interactive Elements Ocr Chemistry Exam Question Answers
- 8. Staying Engaged with Ocr Chemistry Exam Question Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr Chemistry Exam Question Answers
- 9. Balancing eBooks and Physical Books Ocr Chemistry Exam Question Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr Chemistry Exam Question Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ocr Chemistry Exam Question Answers
 - Setting Reading Goals Ocr Chemistry Exam Question Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ocr Chemistry Exam Question Answers
 - Fact-Checking eBook Content of Ocr Chemistry Exam Question Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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