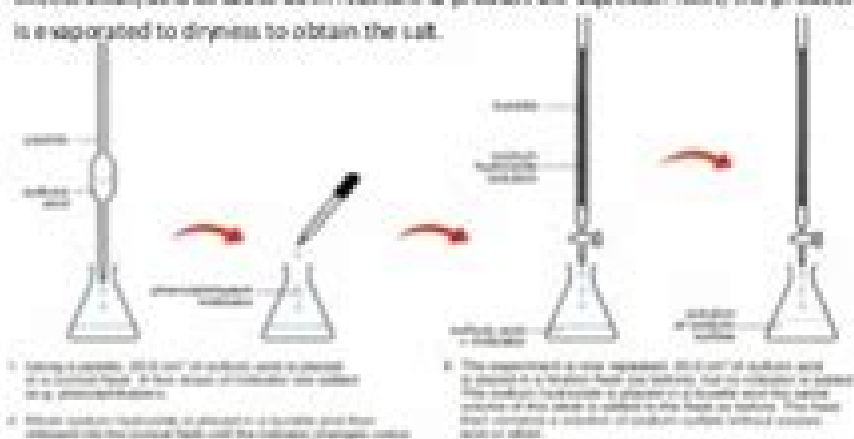


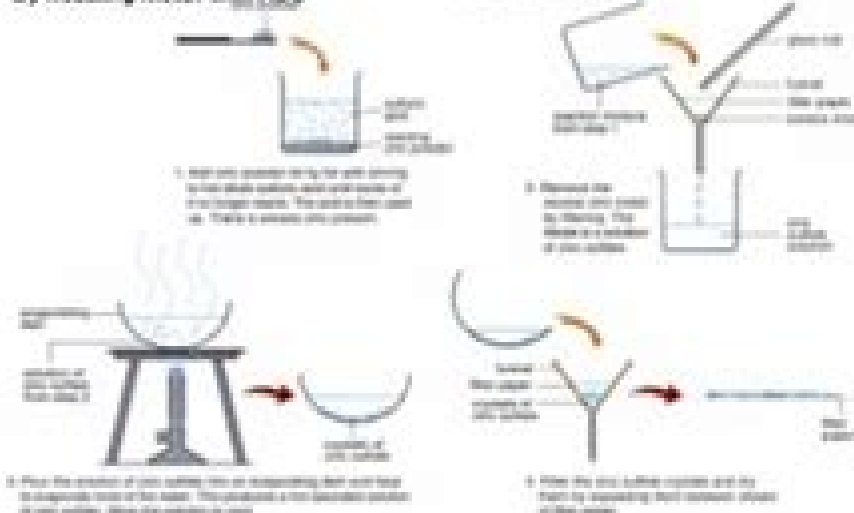
Preparation of Soluble Salts

By Neutralisation

25.0cm³ acid, as standard solution, is placed in conical flask using pipette. Add few drops of indicator & titrate with alkali from burette until indicator changes colour, showing all acid has just reacted. Volume of alkali added is measured. Prepare new 25.0cm³ acid again with pipette & add same volume of alkali as before to prevent excess alkali/acid because both reactant & product are aqueous. Next, the product is evaporated to dryness to obtain the salt.



By Reacting Metal with Acid



Only metals like zinc and magnesium, which moderately react with dilute acids, are used.

E.g. Reacting Zn with H₂SO₄ to prepare ZnSO₄



Zn is added to H₂SO₄ until in excess to ensure no more H₂SO₄ is present. Then the mixture is filtered off to separate Zn from ZnSO₄. The filtrate (ZnSO₄) is then placed in evaporating dish to evaporate most of water then it's cooled after ZnSO₄ crystals are formed. The crystals then filtered and squeezed between filter papers to dry.

By Reacting Insoluble Base with Acid

E.g. Reacting MgO with Acids



The same method as reaction of acid with metal is used, so refer to diagram and above explanation, substituting reactants and products.

By Reacting Carbonate with Acid

E.g. Reacting CaCO₃ with Acids



The same process is used as reaction of acid with metal, just that carbon dioxide is produced. Carbon dioxide can be tested by bubbling it into limewater which will turn limewater colourless to milky.

7.1 Properties and Uses of Ammonia

Ammonia and its Uses

Are produced from nitrogen reacted with hydrogen

For producing: fertilisers, nitric acid, nylon, dyes, cleaners and dry cell

The Manufacture of Ammonia: The Haber Process

The Process

Nitrogen and hydrogen are mixed together in ratio 1:3, where nitrogen is obtained from air and hydrogen is obtained from natural gas, and passed over iron-catalyst.



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Table of Contents O Level Chemistry Notes Studyguide Pk

1. Understanding the eBook O Level Chemistry Notes Studyguide Pk
 - The Rise of Digital Reading O Level Chemistry Notes Studyguide Pk
 - Advantages of eBooks Over Traditional Books
2. Identifying O Level Chemistry Notes Studyguide Pk
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an O Level Chemistry Notes Studyguide Pk
 - User-Friendly Interface
4. Exploring eBook Recommendations from O Level Chemistry Notes Studyguide Pk
 - Personalized Recommendations
 - O Level Chemistry Notes Studyguide Pk User Reviews and Ratings
 - O Level Chemistry Notes Studyguide Pk and Bestseller Lists
5. Accessing O Level Chemistry Notes Studyguide Pk Free and Paid eBooks
 - O Level Chemistry Notes Studyguide Pk Public Domain eBooks
 - O Level Chemistry Notes Studyguide Pk eBook Subscription Services
 - O Level Chemistry Notes Studyguide Pk Budget-Friendly Options
6. Navigating O Level Chemistry Notes Studyguide Pk eBook Formats
 - ePub, PDF, MOBI, and More
 - O Level Chemistry Notes Studyguide Pk Compatibility with Devices
 - O Level Chemistry Notes Studyguide Pk Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of O Level Chemistry Notes Studyguide Pk
 - Highlighting and Note-Taking O Level Chemistry Notes Studyguide Pk
 - Interactive Elements O Level Chemistry Notes Studyguide Pk
8. Staying Engaged with O Level Chemistry Notes Studyguide Pk

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers O Level Chemistry Notes Studyguide Pk
- 9. Balancing eBooks and Physical Books O Level Chemistry Notes Studyguide Pk
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection O Level Chemistry Notes Studyguide Pk
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine O Level Chemistry Notes Studyguide Pk
 - Setting Reading Goals O Level Chemistry Notes Studyguide Pk
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of O Level Chemistry Notes Studyguide Pk
 - Fact-Checking eBook Content of O Level Chemistry Notes Studyguide Pk
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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