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 $2 \text{ CO}_2 \text{ g} + \text{H}_2\text{O g} \rightarrow \text{C}_2\text{H}_2 \text{ g} + \frac{5}{2} \text{O}_2 \text{ g}$ $\Delta h = 2 \text{ h}_2 \text{ g} + \text{C}_2\text{H}_6 \text{ g}$ $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O g} + \text{H}_2 \text{ g} + \frac{1}{2} \text{O}_2 \text{ g} \Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ g} + \frac{7}{2} \text{O}_2 \text{ g} \rightarrow \text{CO}_2 \text{ g} + 3 \text{ H}_2\text{O g}$ $\Delta h = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ g}$

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web chemistry 120 hess s law worksheet calculate h for the reaction $\text{C}_2\text{H}_4 \text{ g} + \text{H}_2 \text{ g} \rightarrow \text{C}_2\text{H}_6 \text{ g}$ from the following data
 $\text{C}_2\text{H}_4 \text{ g} + 3 \text{ O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 2 \text{ H}_2\text{O l}$ $\Delta h = 1411 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ g} + \frac{7}{2} \text{ O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 3 \text{ H}_2\text{O l}$ $\Delta h = 1560 \text{ kJ}$
 $\text{H}_2 \text{ g} + \frac{1}{2} \text{ O}_2 \text{ g} \rightarrow \text{H}_2\text{O l}$ $\Delta h = 285 \text{ kJ}$
 mole

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