

## A2 Unit F334: Chemistry of Materials

### Topic 3: What's in a medicine? (date)

A study of medicines such as aspirin, their development, chemistry and synthesis, illustrating some of the features of the pharmaceutical industry. The chemical ideas in this module are:

- Molecular synthesis: acids, esters, carbonyl compounds
- Acid-base reactions
- Medicines: manufacture and testing
- IR spectroscopy and mass spectroscopy

| Topic                 | Alt code | Number | Assessable learning outcomes   | Number of questions | OSP A2 Revision guide (Page number) | Chemical changes (Page number) | Chemical ideas (Page number) |
|-----------------------|----------|--------|--|---------------------|-------------------------------------|--------------------------------|------------------------------|
| Equilibria            |          | 10401  | Describe acids in terms of the Brønsted-Lowry theory as proton donors, and bases as proton acceptors   | 1                   | 6                                   |                                | 100                          |
|                       |          | 10402  | Identify the proton donor and proton acceptor in an acid-base reaction   | 1                   | 6                                   |                                | 100                          |
| Bonding and structure | 11.1.0   | 10403  | Draw and interpret simple electron-dot-and-cross diagrams to show how atoms (bond through one, covalent and three covalent bonds)  | 2                   | 14                                  |                                | 101, 12                      |
|                       | 11.1.0a  | 10404  | Describe a simple model of metallic bonding  |                     | 1                                   |                                | 41                           |
|                       | 12.1.0   | 10405  | Use the electron-dot-repulsion principle to predict and explain the shapes of simple molecules (such as CH <sub>4</sub> , NH <sub>3</sub> , H <sub>2</sub> O and HF) and ions (such as NH <sub>4</sub> <sup>+</sup> ) with up to six outer pairs of electrons. (any combination of bonding pairs and lone pairs) (no treatment of hybridisation or molecular orbitals is expected) | 1                   | 7                                   |                                | 101, 105                     |

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**Tao Wei**



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