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web solution $v_1 = 250 \text{ l}$ $v_2 = 273 \text{ 30 303 K}$ $t_2 = 273 \text{ 70 343 K}$ we will be applying charles law as pressure is constant here $v_1 t_1 = v_2 t_2$ or $v_2 = v_1 t_1 \times t_2 / 250 \text{ l 303 K} \times 343 \text{ K} / 273 \text{ l}$ at 90°C a helium sample has a volume of 500 ml determine the temperature at which the volume of the liquid will become 240 ml