

4

Atomic Structure
Reviewsheet

A. Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this chapter. Each blank can be completed with a term, short phrase, or number.

Atoms of each element are 1 from the atoms of all other elements. Dalton theorized that atoms are indivisible, but the discovery of 2 particles changed this theory. We now know that atoms are made up of electrons, which have a 3 charge; 4, which have a positive charge; and 5, which are neutral. The latter two particles are found in the 6 of the atom.

It was 7 who discovered the nucleus of the atom. The nucleus has a 8 charge and it occupies a very small volume of the atom. In contrast, the negatively charged 9 occupy most of the volume of the atom.

The number of 10 in the nucleus of the atom is the atomic 11 of that element. Because atoms are electrically neutral, the number of protons and 12 in an atom are equal. The sum of the 13 and neutrons is the mass number. Atoms of the same element are identical in most respects, but they can differ in the number of 14 in the nucleus. Atoms that have the same number of protons but different mass numbers are called 15.

The 16 of an element is the weighted average of the masses of the isotopes of that element. Two isotopes of sulfur are $^{32}_{16}\text{S}$ and $^{34}_{16}\text{S}$. An atom of the sulfur-32 isotope contains 17 protons and 18 neutrons. The sulfur-34 isotope has 19 protons and 20 neutrons.

1. different 4.1
2. sub atomic 4.2
3. negative 4.3
4. protons 4.2
5. neutrons 4.2
6. nucleus 4.3
7. Rutherford 4.3
8. positive 4.3
9. electrons 4.3
10. protons (elec) 4.4
11. number 4.4
12. electrons 4.4
13. protons (elec) 4.5
14. neutrons 4.6
15. isotopes 4.6
16. atomic mass 4.7
17. 16 4.5
18. 16 4.6
19. 16 4.6
20. 18 4.6

Review Sheet Unit 15 Chemistry Answers

Scott C. Dulebohn



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Association. Agricultural Education Section,1961 **Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e** Kenneth W. Raymond,2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond s text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry **Vocational Education Bulletin** United States. Division of Vocational Education,1961

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