

CLO- Classification of Elements & Periodicity in Properties

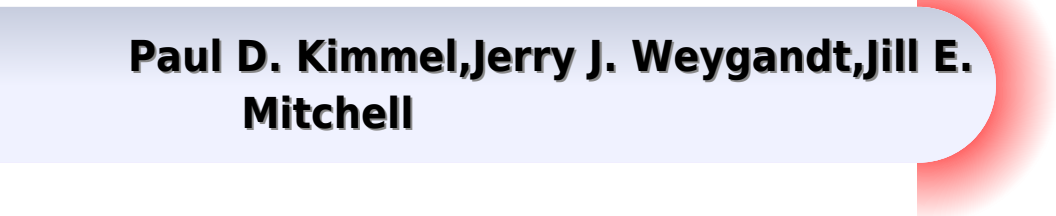
MULTIPLE CHOICE QUESTIONS:

Choose the one correct option in the following:

- Identify the least stable ion amongst the following :
(a) Li^+ (b) Be^+ (c) B^+ (d) C^+
- The correct order of radii is :
(a) $\text{N} < \text{Be} < \text{B}$ (b) $\text{F}^- < \text{O}^{2-} < \text{N}^{3-}$ (c) $\text{N} < \text{Li} < \text{K}$ (d) $\text{Fe}^{2+} < \text{Fe}^{3+} < \text{Fe}^{4+}$
- Which one of the following has largest size :
(a) Na (b) Na^+ (c) Na^- (d) Li^+
- The number of elements in the 4th period of periodic table is :
(a) 8 (b) 10 (c) 18 (d) 32
- Which of the following has least electron affinity (electron gain enthalpy) :
(a) O (b) N (c) Ar (d) F
- An element of atomic number 29 belongs to which of the following block of the periodic table :
(a) s-block (b) p-block (c) d-block (d) f-block
- Which one of the following arrangements not truly represents the property indicated against it:
(a) $\text{Br}_2 < \text{Cl}_2 < \text{F}_2$ – electronegativity (b) $\text{Br}_2 < \text{F}_2 < \text{Cl}_2$ – Electron Affinity
(c) $\text{Br}_2 < \text{Cl}_2 < \text{F}_2$ – Bond energy (d) $\text{Br}_2 < \text{Cl}_2 < \text{F}_2$ – Oxidising Power
- Alkali metals in each period have :
(a) Smallest size (b) Lowest IE (c) Highest IE (d) Highest EN
- A sudden large jump between the value of second and third ionization energies of elements would be associated with which of the following electronic configuration :
(a) $2\text{p}^4 3\text{s}^1$ (b) $3\text{s}^2 3\text{p}^1$ (c) $3\text{s}^2 3\text{p}^2$ (d) $2\text{p}^4 3\text{s}^2$
- Which of the following pairs are chemically dissimilar :
(a) Na and K (b) Ba and Sr (c) Zr and Hf (d) Ca and Zn
- Collective name given to the elements with outer shell configuration $ns^2 np^6$ is :
(a) Chalkogens (b) Alkaline earths metals (c) Transition metals (d) Nobel gases
- Modern extended periodic table was given by :
(a) Bohr (b) Lothar mayer (c) Moseley (d) None alone
- The most electronegative elements is :
(a) Hydrogen (b) Fluorine (c) Chlorine (d) Oxygen
- The element having electronic configuration :
 $[\text{Kr}]4\text{d}^{10}4\text{f}^{14}5\text{s}^2 5\text{p}^6 5\text{d}^1 6\text{s}^2$ belongs to
(a) s-block (b) p-block (c) d-block (d) f-block
- An elements R forms the highest oxide R_2O_5 . R belong to :

Periodicity Multiple Choice Questions

**Paul D. Kimmel, Jerry J. Weygandt, Jill E.
Mitchell**



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This book delves into Periodicity Multiple Choice Questions. Periodicity Multiple Choice Questions is an essential topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Periodicity Multiple Choice Questions, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Periodicity Multiple Choice Questions
- Chapter 2: Essential Elements of Periodicity Multiple Choice Questions
- Chapter 3: Periodicity Multiple Choice Questions in Everyday Life
- Chapter 4: Periodicity Multiple Choice Questions in Specific Contexts
- Chapter 5: Conclusion

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4. In chapter 3, this book will examine the practical applications of Periodicity Multiple Choice Questions in daily life. This chapter will showcase real-world examples of how Periodicity Multiple Choice Questions can be effectively utilized in everyday scenarios.

5. In chapter 4, the author will scrutinize the relevance of Periodicity Multiple Choice Questions in specific contexts. The fourth chapter will explore how Periodicity Multiple Choice Questions is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, the author will draw a conclusion about Periodicity Multiple Choice Questions. This chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Periodicity Multiple Choice Questions.

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