Chapter 4 - Practice Problems Solutions

 Complete and balance the following reactions. Remember to form the stable product and then balance. Assume the product is solid. What type of reactions are these?

____ Mg (s) + ____ O₂ (g)
$$\rightarrow$$
 ____ Na (s) + ____ Cl₂ (g) \rightarrow 2 Mg (s) + O₂ (g) \rightarrow 2 Na (s) + Cl₂ (g) \rightarrow 2 Na Cl (s)

____ Al (s) + ____ O₂ (g)
$$\rightarrow$$
 ____ Ca (s) + ____ Br₂ (g) \rightarrow 4 Al (s) + 3 O₂ (g) \rightarrow 2 Al₂O₃ (s) Ca (s) + Br₂ (g) \rightarrow CaBr₂ (s)

These are all combination reactions.

2) Balance the following reactions. What type of reactions are these?

$$C_3H_8(g) + ___O_2(g) \Rightarrow __CO_2(g) + ___H_2O(g)$$

 $C_3H_8(g) + 5 O_2(g) \Rightarrow __CO_2(g) + 4 H_2O(g)$
 $C_4H_8(g) + __O_2(g) \Rightarrow __CO_2(g) + __H_2O(g)$
 $C_4H_8(g) + 15 O_2(g) \Rightarrow 12 CO_2(g) + 6 H_2O(g)$

These are both combustion reactions.

3) Balance the following reactions. What type of reactions are these?

$$H_2O(I)$$
 → _____ $H_2(g)$ + _____ $O_2(g)$ _____ $KCIO_3(s)$ → _____ $KCI(s)$ + _____ $O_2(g)$
2 $H_2O(I)$ → 2 $H_2(g)$ + $O_2(g)$ 2 $KCIO_3(s)$ → 2 $KCI(s)$ + 3 $O_2(g)$

These are both decomposition reactions.

4) Balance the following reactions. What type of reactions are these?

____ Mg (s) + ____ HCl (aq)
$$\rightarrow$$
 ____ MgCl₂ (aq) + ____ H₂ (g)
Mg (s) + 2 HCl (aq) \rightarrow MgCl₂ (aq) + H₂ (g)
____ Ag (s) + ____ H₂S (aq) \rightarrow ____ Ag₂S (s) + ____ H₂ (g)
2 Ag (s) + H₃S (aq) \rightarrow Ag₂S (s) + H₂ (g)

These are both single displacement reactions.

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Table of Contents Sample Problem Solutions From Chapter 4 Csus

- 1. Understanding the eBook Sample Problem Solutions From Chapter 4 Csus
 - The Rise of Digital Reading Sample Problem Solutions From Chapter 4 Csus
 - o Advantages of eBooks Over Traditional Books
- 2. Identifying Sample Problem Solutions From Chapter 4 Csus
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Sample Problem Solutions From Chapter 4 Csus
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Sample Problem Solutions From Chapter 4 Csus

- Personalized Recommendations
- Sample Problem Solutions From Chapter 4 Csus User Reviews and Ratings
- Sample Problem Solutions From Chapter 4 Csus and Bestseller Lists
- 5. Accessing Sample Problem Solutions From Chapter 4 Csus Free and Paid eBooks
 - Sample Problem Solutions From Chapter 4 Csus Public Domain eBooks
 - Sample Problem Solutions From Chapter 4 Csus eBook Subscription Services
 - Sample Problem Solutions From Chapter 4 Csus Budget-Friendly Options
- 6. Navigating Sample Problem Solutions From Chapter 4 Csus eBook Formats
 - o ePub, PDF, MOBI, and More
 - Sample Problem Solutions From Chapter 4 Csus Compatibility with Devices
 - Sample Problem Solutions From Chapter 4 Csus Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Sample Problem Solutions From Chapter 4 Csus
 - Highlighting and Note-Taking Sample Problem Solutions From Chapter 4 Csus
 - Interactive Elements Sample Problem Solutions From Chapter 4 Csus
- 8. Staying Engaged with Sample Problem Solutions From Chapter 4 Csus
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - $\circ\,$ Following Authors and Publishers Sample Problem Solutions From Chapter 4 Csus
- 9. Balancing eBooks and Physical Books Sample Problem Solutions From Chapter 4 Csus
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Sample Problem Solutions From Chapter 4 Csus
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Sample Problem Solutions From Chapter 4 Csus
 - Setting Reading Goals Sample Problem Solutions From Chapter 4 Csus
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Sample Problem Solutions From Chapter 4 Csus

- Fact-Checking eBook Content of Sample Problem Solutions From Chapter 4 Csus
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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