

CHAPTER 9 REVIEW

Stoichiometry

KEY

SECTION 3

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.

$$\% \text{ yield} = \frac{22 \text{ g}}{25 \text{ g}} \times 100 = 88\%$$

2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation:

2.0 mol x 3 N_2

ratio

6.0 mol N_2 12.0 mol H_2

- a. Which chemical is in excess? What is the excess in moles?

$$\text{N}_2: \left(\frac{2 \text{ mol}}{1 \text{ H}_2} \right) \left(\frac{1 \text{ N}_2}{3 \text{ H}_2} \right) = 9.0 \text{ mol}$$

not $3 \times \text{N}_2 \Rightarrow \text{H}_2$ is limiting8.0 mol NH_3

- b. Theoretically, how many moles of NH_3 will be produced?

$$\text{mol NH}_3 = \left(\frac{12.0 \text{ mol}}{\text{H}_2} \right) \left(\frac{2 \text{ NH}_3}{3 \text{ H}_2} \right) = 8.0 \text{ mol}$$

6.4 mol NH_3

- c. If the percentage yield of NH_3 is 80%, how many moles of NH_3 are actually produced?

$$\% \text{ yield} = \frac{\text{actual}}{\text{theor}} \times 100$$

$$\text{actual} = \frac{(\% \text{ yield})(\text{theor})}{100} = \frac{(80)(8.0)}{100} = 6.4 \text{ mol}$$

3. 0.050 mol of $\text{Ca}(\text{OH})_2$ are combined with 0.080 mol of HCl according to the following equation:

0.10 mol HCl

- a. How many moles of HCl are required to neutralize all 0.050 mol of $\text{Ca}(\text{OH})_2$?

$$\text{mol HCl} = \left(\frac{0.050 \text{ mol}}{\text{Ca}(\text{OH})_2} \right) \left(\frac{2 \text{ HCl}}{1 \text{ Ca}(\text{OH})_2} \right) = 0.10 \text{ mol HCl}$$

limiting

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