

Fill in the blank in each sentence and/or write the word.

1. The theory that explains the behavior of gases on the molecular level is called the _____, which is based on assumptions related to the motion of particles of matter.
2. Use the kinetic theory to explain why a balloon filled with air expands when it is brought inside a warm room from the cold outdoors.
 According to the kinetic theory, particles have more kinetic energy when they are in a warmer environment. The particles move faster and hit the walls of the balloon more often, causing it to expand.
3. Define "ideal gas."
4. Gases deviate most from ideal behavior under conditions of _____ temperature and _____ pressure.
5. Pressure is defined as _____ per unit _____. The earth's atmosphere has weight, which creates _____.
6. Standard temperature has been established as _____ °C or _____ °F. Standard pressure has been established as _____ atm, _____ mm Hg, or _____ kPa. The atmosphere has a standard temperature and pressure of _____.
7. A _____ is used to measure atmospheric pressure, which is _____, or total atmospheric gas pressure.
8. When atmospheric pressure increases, how does the height of the column of mercury change?
9. At constant _____, the volume of a sample of gas _____ as temperature increases.
10. At constant pressure, the volume of a sample of gas is _____ as temperature decreases.

Review Sheet Unit 9 Chemistry Answers

S Ben Porath



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