

Fill in the gaps in statements and/or equations.

1. The theory that explains the behaviour of gases on the molecular level is called the kinetic theory, which is based on assumptions related to the ideal gas of non-interacting particles.
2. Use the kinetic theory to explain why a balloon filled with air expands when it is brought inside a warm room from the cold outdoors.
 According to the kinetic theory, particles have more kinetic energy (temperature) inside a warm room than outside. The particles will move faster and exert more pressure on the walls of the balloon, causing it to expand.
3. Define "1 bar" and "1 atm".
 1 bar is defined as 10⁵ Pa and 1 atm is defined as 1.01325 × 10⁵ Pa.
4. Gases deviate most from ideal behaviour under conditions of high pressure and low temperature.
5. Pressure is defined as force per unit area. The earth's atmosphere has weight, which creates atmospheric pressure.
6. Standard temperature has been established as 0 °C or 273.15 K. Standard pressure has been established as 1 bar or 10⁵ Pa. The standard unit for pressure is Pa. The atmosphere has a standard temperature and pressure of 15 °C and 1.01325 × 10⁵ Pa.
7. A manometer is used to measure atmospheric pressure, which is relative to the local atmospheric pressure.
8. Which atmospheric pressure increases, sea level or the height of a mountain? Sea level.
9. At constant temperature, the volume of a sample of gas decreases as temperature increases.
10. At constant pressure, the volume of a sample of gas increases as temperature increases.

Review Sheet Unit 9 Answer Key

K Morrison



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