

Graphing Periodic Trends Activity

Name: _____ Period: _____

Use the information in these tables to complete the graph as described below

Symbol	Atomic Radius (Picometers)	First Ionization energy (kilojoules/mole)	Electronegativity (4-point scale)
H	31	1312	2.1
Li	128	520	1.0
Na	166	496	0.9
K	203	410	0.8
Rb	220	403	0.8
Cs	244	376	0.7

Symbol	Atomic Radius (Picometers)	First Ionization energy (kilojoules/mole)	Electronegativity (4-point scale)
Na	166	496	0.9
Mg	141	738	1.2
Al	121	578	1.5
Si	111	787	1.8
P	107	1012	2.1
S	105	1000	2.5
Cl	102	1251	3.0
Ar	106	1521	---

1. Using colored pencils or pen, list under the symbol (in this order!) the Atomic Radius, First Ionization Energy, and Electronegativity.
2. Use a different color for each property. Example: Write all of the Atomic radius values in red, all of the First ionization energies in green, and all of the Electronegativities in blue. Of course, you can pick any colors that are available, as long as you are consistent.
3. Observe the trends in each property as you go down the Alkali metal group, and as you go across Period 3.
4. Write out each of the statements written on the opposite side, completing each statement with the observed trend (increase or decrease)

H	Across a period (→) atomic radius tends to _____. Across a period (→) first ionization energy tends to _____. Across a period (→) electronegativity tends to _____.						
Li							
Na 166 496 0.9	Mg	Al	Si	P	S	Cl	Ar
K	Down a Group (↓) atomic radius tends to _____. Down a Group (↓) first ionization energy tends to _____. Down a Group (↓) electronegativity tends to _____.						
Rb							
Cs							

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web magnesium carbonate nitric acid magnesium nitrate water carbon dioxide one mole of magnesium carbonate MgCO_3 and two moles of nitric acid HNO_3 react to form one mole of magnesium nitrate $\text{Mg}(\text{NO}_3)_2$ one mole of water H_2O and one mole of carbon dioxide CO_2 show chemical structure image reactants

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web magnesium is already balanced on the reactant and product side balancing nitrogen by multiplying HNO_3 by 2 we get Mg s magnesium 2 HNO_3 l nitric acid $\text{Mg}(\text{NO}_3)_2$ s magnesium nitrate H_2 g hydrogen gas oxygen and hydrogen get automatically balanced by balancing nitrogen

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web magnesium nitrate sulfuric acid magnesium sulfate nitric acid

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web the balanced net ionic equation for the reaction of magnesium carbonate with nitric acid is a MgCO_3 s 2H^+ aq Mg^{2+} aq CO_2 g H_2O l b Mg^{2+} aq CO_3^{2-} aq 2H^+ aq 2NO_3^- aq Mg write the balanced equation when nitric acid reacts with calcium hydroxide to produce water and calcium nitrate

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web magnesium oxide nitric acid gives rise to a magnesium phosphate hydrogen balanced equation on may 27 2019 how do you balance magnesium and nitric acid you ll need to count all of atoms on both directions of the chemical equation to balance Mg HNO_3 mm NO_3 2 NO_2 H_2O

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web how to balance HNO_3 Mg $\text{Mg}(\text{NO}_3)_2$ H_2 word equation nitric acid magnesium magnesium nitrate hydrogen gas type of chemical reaction for this reaction we have a single displacement reaction balancing strategies in this single displacement reaction the Mg replaces the H in the HNO_3

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web nov 20 2020 in this video we ll balance the equation Mg HNO_3 $\text{Mg}(\text{NO}_3)_2$ NO_2 H_2O and provide the correct coefficients for each compound to balance Mg HNO_3 $\text{Mg}(\text{NO}_3)_2$ NO_2 H_2O you ll need to be

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Periodic Trends Extension Questions Answer Key

web magnesium hydroxide nitric acid balanced total equation $\text{Mg(OH)}_2(\text{aq}) + 2\text{HNO}_3(\text{aq}) \rightarrow \text{Mg(NO}_3)_2(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$ total ionic equation
 $\text{Mg}^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq}) + 2\text{H}^{+}(\text{aq}) + 2\text{NO}_3^{-}(\text{aq}) \rightarrow \text{Mg}^{2+}(\text{aq}) + 2\text{NO}_3^{-}(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$ net ionic equation $2\text{OH}^{-}(\text{aq}) + 2\text{H}^{+}(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l})$ sulfuric acid and lithium hydroxide
balanced total equation $\text{H}_2\text{SO}_4(\text{aq}) + 2\text{LiOH}(\text{aq}) \rightarrow \text{Li}_2\text{SO}_4(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$