Study Guide

Key Concepts

(14.1 Properties of Gases

- Gases are easily compressed because of the space between particles in a gas.
- The amount of gas (n), volume (V), and temperature (T) are factors that affect gas pressure (P).

C 14.2 The Gas Laws

- · As the pressure of a gas increases, the volume decreases. If the temperature is constant.
- · As the temperature of an enclosed gas increases, the volume increases, if the pressure is constant.
- · As the temperature of an enclosed gas increases, the pressure increases, if the volume is constant.

 The combined gas law allows you to do calculations for situations in which only the amount of gas is constant.

C 14.3 Ideal Gases

- · To calculate the number of moles of a contained gas requires an expression that contains the variable 12.
- · Real gases differ most from an ideal gas at low temperatures and high pressures.

14.4 Gases: Mixtures and Movements

- . In a mixture of gases, the total pressure is the sum of the partial pressures of the gases.
- Gases of lower molar mass diffuse and effuse faster than gases of higher molar mass.

Vocabulary

- Boyle's law (p. 418).
- Charles's law (p. 420)
- combined gas law (p. 424).
- compressibility (p. 413)

- Dalton's law of partial pressures (p. 432)
- diffusion (p. 435).
- effusion (p. 435).
- Gay-Lussac's law (p. 422).

- Graham's law of effusion. (p.435)
- ideal gas constant (p. 426)
- ideal gas law (p. 426)
- partial pressure (p. 432).

Key Equations

- Boyle's law: P₁ × V₁ = P₂ × V₂
- Charles's law: V = V2
- Gay-Lussac's law: P₁ = P₂ / T.
- Combined gas law: $\frac{P_1 \times V_1}{T_1} = \frac{P_2 \times V_2}{T_2}$

- Ideal gas law: P × V = n × R × T or PV = nRT
- Dalton's law: P_{need} = P₁ + P₂ + P₃ + ...
- Graham's law: $\frac{Rate_A}{Rate_B} = \sqrt{\frac{molar \ mass_B}{molar \ mass_A}}$

Organizing Information

Use these terms to construct a concept. map that organizes the major ideas of this chapter.



Concept Map 14 Solve the Concept Map with the help of an interactive guided tutorial.

with Chemiasap

Boyle's Law

Charles's law.

partial pressure

combined gas law.

diffusion

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William A. Lowenbach

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Table of Contents Section Assessment Prentice Hall Chemistry

- 1. Understanding the eBook Section Assessment Prentice Hall Chemistry
 - The Rise of Digital Reading Section Assessment Prentice Hall Chemistry
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Section Assessment Prentice Hall Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section Assessment Prentice Hall Chemistry
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Section Assessment Prentice Hall Chemistry
 - Personalized Recommendations
 - Section Assessment Prentice Hall Chemistry User Reviews and Ratings
 - Section Assessment Prentice Hall Chemistry and Bestseller Lists

- 5. Accessing Section Assessment Prentice Hall Chemistry Free and Paid eBooks
 - Section Assessment Prentice Hall Chemistry Public Domain eBooks
 - Section Assessment Prentice Hall Chemistry eBook Subscription Services
 - Section Assessment Prentice Hall Chemistry Budget-Friendly Options
- 6. Navigating Section Assessment Prentice Hall Chemistry eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - Section Assessment Prentice Hall Chemistry Compatibility with Devices
 - Section Assessment Prentice Hall Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section Assessment Prentice Hall Chemistry
 - Highlighting and Note-Taking Section Assessment Prentice Hall Chemistry
 - Interactive Elements Section Assessment Prentice Hall Chemistry
- 8. Staying Engaged with Section Assessment Prentice Hall Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section Assessment Prentice Hall Chemistry
- 9. Balancing eBooks and Physical Books Section Assessment Prentice Hall Chemistry
 - \circ Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Section Assessment Prentice Hall Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Section Assessment Prentice Hall Chemistry
 - Setting Reading Goals Section Assessment Prentice Hall Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section Assessment Prentice Hall Chemistry
 - Fact-Checking eBook Content of Section Assessment Prentice Hall Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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