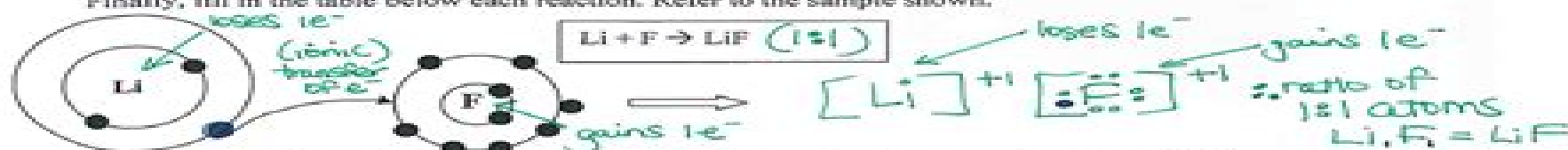


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### Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



| Atoms | Valence electrons | Electron transfer from/to each atom | Ions formed in the product |
|-------|-------------------|-------------------------------------|----------------------------|
| Li    | 1 ve              | transfers $1e^-$ to F               | $Li^+$                     |
| F     | 7 ve              | accepts $1e^-$ from Li              | $F^-$                      |

| Reactions                                | Atoms | Valence electrons | Electron transfer from/to each atom | Ions formed in the product |
|------------------------------------------|-------|-------------------|-------------------------------------|----------------------------|
| 1) $Li + Cl \Rightarrow LiCl$            | Li    | 1 ve              | donates $1e^-$                      | $Li^+$                     |
| $Li + Cl \rightarrow [Li]^+ [Cl:]^{-1}$  | Cl    | 1 ve              | accepts $1e^-$                      | $Cl^-$                     |
| 2) $Ca + O \Rightarrow CaO$              | Ca    | 2 ve              | donates $2e^-$                      | $Ca^{+2}$                  |
| $Ca + O \rightarrow [Ca]^{+2} [O:]^{-2}$ | O     | 6 ve              | accepts $2e^-$                      | $O^{-2}$                   |
| 3) $Be + F \Rightarrow BeF_2$            | Be    | 2 ve              | donates $2e^-$ to each F            | $Be^{+2}$                  |
| $Be + F \rightarrow [Be]^{+2} [F:]^{-1}$ | F     | 7 ve              | each F accepts $1e^-$ each          | $F^-$                      |
| 4) $Mg + S \Rightarrow MgS$              | Mg    | 2 ve              | donates $2e^-$ to S                 | $Mg^{+2}$                  |
| $Mg + S \rightarrow [Mg]^{+2} [S:]^{-2}$ | S     | 6 ve              | accepts $2e^-$                      | $S^{-2}$                   |
| 5) $K + F \Rightarrow KF$                | K     | 1 ve              | donates $1e^-$                      | $K^+$                      |
| $K + F \rightarrow [K]^+ [F:]^{-1}$      | F     | 7 ve              | accepts $1e^-$                      | $F^-$                      |

# Section 6 Ionic Bonding Worksheet Answer Key

**Marcel A. Müller**



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