

Chapter 1 Answers

Practice 1-1

1. $7 + x$ 2. $3p$ 3. $10 - m$ 4. $7 - n$
 5. $2g$ 6. $m + 3$ 7. 8 divided by a 8. 10 less than x
 9. 13 more than x 10. 2 more than the product of a and b
 11. $8 - n$ 12. $4 + a$ 13. $2a$ 14. $3 + n$ 15. $10 + \frac{a}{13}$
 16. $n - 12$ 17. $c = 24.95s$ 18. $g = 30r$ 19. $m = 0.10df$
 20. $n = 48 - g$ 21. $c = 8p$ 22. $c = 32.95p$
 23–28. Choice of variables may vary.
 t = number of tickets, c = total cost;
 h = number of hours, d = distance;
 h = number of hours, p = total pay;
 c = change, a = total cost;
 d = number of days, l = length;
 r = miles traveled, t = miles remaining
 23. $c = 3.50r$ 24. $d = 55h$ 25. $p = 5h$ 26. $a = 10 - c$
 27. $l = 0.45d$ 28. $r = 500 - t$

Practice 1-2

1. 52 2. 2 3. 38 4. 4 5. 18.9 6. 87 7. 25 8. 7 9. 2
 10. 1 11. 9 12. 30.4 13. 19 14. 5 15. 2 16. 6 17. 26
 18. 29 19. 31 20. 88 21. 14.18 22. 85 23. 28 24. 12
 25. 26 26. 5 27. 55 28. 56 29. 72 30. 60 31. 40 32. 2
 33. 11 34. 4 35. 131 36. 141 37. 33 38. 16 39. 22
 40. 12 41. 2 42. 1 43. 56 44. \$33.95 45. \$63.34 46. 24
 47. 2.25 48. 0 49. 99 50. 5 51. 6

Practice 1-3

1. rational, real 2. rational, real 3. natural, whole, integers, rational, real 4. irrational, real 5. irrational, real
 6. integers, rational, real 7. rational, real 8. natural, whole, integers, rational, real 9. true 10. false; -2 11. true
 12. false; $7 \times 2 = 14$ 13. $>$ 14. $<$ 15. $<$ 16. $=$ 17. $>$
 18. $<$ 19. $-\frac{8}{9}$, $-\frac{22}{25}$, $-\frac{7}{8}$ 20. $-3\frac{12}{25}$, -3.45 , $-3\frac{4}{9}$
 21. $-\frac{1}{3}$, $-\frac{1}{4}$, $-\frac{1}{2}$ 22. $-1\frac{7}{9}$, $-1\frac{3}{4}$, -1.7 23. $-\frac{7}{8}$, $-\frac{3}{4}$, $-\frac{2}{3}$
 24. $2\frac{4}{8}$, 2.7 , $2\frac{3}{4}$ 25. whole 26. rational 27. whole
 28. irrational 29. $\frac{3}{10}$ 30. 327 31. 3.46 32. $\frac{1}{2}$
 33. \$154.23, 0.0375, $\frac{30}{325}$, \$.48, 0.055, and $3\frac{1}{2}$ are rational and real; \$8000 and \$1540 are whole, integers, rational, and real.

Practice 1-4

1. $p = 9t$ 2. $c = 13n$ 3. $m = 5n + 20$
 4. $c = 2.50w + 7.50$ 5. number of hours; amount earned
 6. number of people; cost 7. number of minutes; volume
 8. price; number of people 9. 96; 120; $b = 24c$
 10. \$.90; \$1.00; $c = 0.10m + 0.50$

Practice 1-5

1.



2.



3. positive correlation 4. negative correlation
 5. no correlation 6. no correlation 7. positive correlation
 8. no correlation; a person's age is not related to the number of pets he/she may have. 9. negative correlation; the more times you brush, the fewer the cavities you are likely to have.
 10. positive correlation; the more rainy days, the more umbrellas are likely to be sold. 11a. Each point represents one infant's age and weight. 11b. There will be points with the same y -coordinate. 11c. The points will have the same x -coordinate. 11d. Yes, there is a positive correlation. Weight increases with age.

Practice 1-6

1. 2.5; 2; 1; median 2. \$120.50; \$125; no mode; mean
 3. \$31.16; \$31.25; no mode; median 4. 7.11; 8; 8 and 10; median
 5. \$2.81 6. 3
 7. $\frac{4.8 + 1.6 + 5.2 + x}{4} = 3.7$; 3.2
 8. $\frac{40 + 98 + 94 + 102 + 21 + x}{6} = 88$; 173
 9. $\frac{100 + 172 + 85 + 92 + x}{5} = 115$; 126
 10. $\frac{25.6 + 19.3 + 19 + x}{4} = 24$; 32.1
 11. yes; The range is 0.9 fl oz. 12. no; The mean was 6.7 ppm.
 13. 98

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Section 36 1 Answer Key

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web the purpose of the copper iron stoichiometry lab was to determine which of the two possible iron sulfate compounds

formed when reacting aqueous copper ii sulfate with solid iron metal the techniques used in this lab were weighing by

difference quantitative transfer and vacuum filtration

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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron

metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a

byproduct of the process as iron ions the iron metal dissolves

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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and percent yield of this reaction will be calculated

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web purpose in this experiment you will observe a single replacement reaction you will also use stoichiometry to predict what the theoretical yield of product is and calculate a percent yield reaction iron copper ii sulfate copper iron ii sulfate pre lab questions write the balanced equation

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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate CuSO_4 and 2 0153 grams of iron powder Fe the goal of this experiment was to determine the product of copper ii sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web stoichiometry lab the reaction of iron with copper ii sulfate the study of stoichiometry deals with the calculation of quantities in a chemical reaction how much product will be produced how much reactant do you need to make that much product

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between

metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web stoichiometric principles can be used to determine the chemical equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction the iron ions dissolve in the copper ii sulfate solution to produce copper ii ions

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web under the finalise of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away copper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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web stoichiometry lab iron and copper sulfate sample data data and observations mass of copper ii sulfate pentahydrate 3.27 g of g mass of filter paper mass of filter paper and dry copper mass of dry copper 1.28 g