

1. Solve for unknown x using the quadratic formula.

$$ax^2 + bx + c = 0$$

$$d = \frac{(x+1)(x+2)=0}{x=-1 \text{ OR } x=-2} \quad b=3 \quad c=2$$

$$1x^2 + 3x + 2 = 0$$

$$x = \frac{-b \pm \sqrt{b^2 - 4ac}}{2a}$$

$$= \frac{-3 \pm \sqrt{3^2 - 4(1)(2)}}{2(1)}$$

$$= \frac{-3 \pm \sqrt{9-8}}{2}$$

$$= \frac{-3 \pm \sqrt{1}}{2} = \frac{-3 \pm 1}{2}$$

$$x = \frac{-3+1}{2} = \frac{-2}{2} = -1 \quad \text{OR} \quad x = \frac{-3-1}{2} = \frac{-4}{2} = -2$$

$$1x^2 - x - 2 = 0 \quad x=1 \text{ OR } x=-2$$

$$(-1)^2 - 4(1)(-2) = 1 + 8 = 9, \quad \sqrt{9} = 3$$

$$x = \frac{-(-1) \pm 3}{2(1)} = \frac{1 \pm 3}{2}$$

$$x = \frac{1+3}{2} = 2 \quad \text{OR} \quad x = \frac{1-3}{2} = -1$$

$$(x-2)(x+1) = 0$$

$$x = \frac{-b \pm \sqrt{b^2 - 4ac}}{2a}$$

$$a=1 \quad b=-3 \quad c=2$$

$$1x^2 - 3x + 2 = 0$$

$$b^2 - 4ac = (-3)^2 - 4(1)(2) = 9 - 8 = 1, \quad \sqrt{1} = 1$$

$$x = \frac{-(-3) \pm 1}{2(1)} = \frac{3 \pm 1}{2}$$

$$x = \frac{3+1}{2} = 2 \quad \text{OR} \quad x = \frac{3-1}{2} = 1$$

$$(x-2)(x-1)=0$$

$$x=2 \text{ OR } x=1$$

$$10x^2 + 10x - 20 = 0$$

$$10^2 - 4(10)(-20) = 100 + 800 = 900, \quad \sqrt{900} = 30$$

$$x = \frac{-10 \pm 30}{2(10)} = \frac{-10 \pm 30}{20}$$

$$x = \frac{-10+30}{20} = \frac{20}{20} = 1 \quad \text{OR}$$

$$x = \frac{-10-30}{20} = \frac{-40}{20} = -2$$

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g Al_2O_3 $\times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}_2\text{O}_3$ c determine the percent yield

yield actual $\times 100 = 260.0 \text{ g} \times 100 / 86.01 \text{ theoretical} = 302.3 \text{ g}$

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web limiting reagents and percentage yield worksheet 1 consider the reaction $\text{I}_2\text{O}_5 + 5\text{CO} \rightarrow 5\text{CO}_2 + \text{I}_2$ a 80.0 grams of iodine v oxide I_2O_5 reacts with 28.0 grams of carbon monoxide CO determine the mass of iodine I_2 which could be produced
 80 g I_2O_5 1 mol I_2O_5 1 mol I_2 $\times$ 1 333.8 g I_2O_5 1 mol I_2O_5 28 g CO 1 mol CO

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web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage
 text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3