

- _____ 1. Which reaction represents natural nuclear decay?
- A) $\text{KClO}_3 \rightarrow \text{K}^+ + \text{ClO}_3^-$
B) ${}_{92}^{235}\text{U} \rightarrow {}_2^4\text{He} + {}_{90}^{231}\text{Th}$
C) $\frac{1}{7}{}^4\text{N} + \frac{4}{2}\text{He} \rightarrow \frac{1}{8}{}^7\text{O} + \frac{1}{1}\text{H}$
D) $\text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O}$
- _____ 2. Positrons and beta particles have
- A) different charges and different masses
B) the same charge and different masses
C) the same charge and the same mass
D) different charges and the same mass
- _____ 3. As a radioactive element emits gamma radiation only, the atomic number of the element
- A) decreases
B) increases
C) remains the same
- _____ 4. Which balanced equation represents a spontaneous radioactive decay?
- A) $\text{H}_2\text{CO}_3 \rightarrow \text{H}_2\text{O} + \text{CO}_2$
B) $\frac{1}{7}{}^4\text{N} + \frac{1}{0}\text{n} \rightarrow \frac{1}{6}{}^4\text{C} + \frac{1}{1}\text{P}$
C) $\frac{14}{6}\text{C} \rightarrow \frac{14}{7}\text{N} + \frac{0}{-1}\text{e}$
D) $14\text{C} + \text{Ca}_3(\text{PO}_4)_2 \rightarrow 3\text{CaC}_2 + 2\text{P} + 8\text{CO}$

- _____ 5. Given the reaction:
- $${}_{88}^{226}\text{Ra} \rightarrow {}_{86}^{222}\text{Rn} + X$$
- Which type of emanation is represented by X ?
- A) positron B) alpha particle
C) proton D) beta particle
- _____ 6. Given the nuclear equation:
- $${}_{10}^{19}\text{Ne} \rightarrow X + {}_9^{19}\text{F}$$
- What particle is represented by X ?
- A) neutron B) alpha
C) positron D) beta
- _____ 7. In the reaction ${}_{93}^{239}\text{Np} \rightarrow {}_{94}^{239}\text{Pu} + X$, what does X represent?
- A) an alpha particle B) a beta particle
C) a neutron D) a proton

Section Quiz Nuclear Decay Study Guide

Jessica J Manson



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