

Name _____ Period _____

Acids and Bases Reinforcement Activity

Write "A" on the line if the following substance is an acid. Write "B" on the line if it is a base. If the substance is generally not an acid or a base, write "N" for neither. Most bases are metal hydroxides (OH), metal oxides (O^{2-}), or ammonium (NH_4^+) hydroxides/oxides.

1. H_2O _____

2. H_2SO_4 _____

3. $NaOH$ _____

4. MgO _____

5. HCN _____

6. $HClO_4$ _____

7. $HClO_3$ _____

8. $HClO_2$ _____

9. SO_3^{2-} _____

10. $HClO$ _____

11. $NaOH$ _____

12. $LiOH$ _____

13. $Ba(OH)_2$ _____

14. NH_4OH _____

15. $Ca(OH)_2$ _____

16. CaO _____

17. HNO_3 _____

18. $C_6H_5CO_2H$ _____

19. HNO_2 _____

20. HCO_2CO_2H _____

21. PO_4^{3-} _____

22. Mg^{2+} _____

23. $H_3C_7O_5$ _____

24. H_3BO_3 _____

25. H_2SO_3 _____

26. SeO _____

27. $RbOH$ _____

28. Rb_2O _____

29. HCl _____

30. Li_2O _____

31. HBr _____

32. Na_2O _____

33. HI _____

34. HF _____

35. K_2O _____

36. H_2S _____

37. C_2H_6 _____

38. C_3H_8 _____

39. $KMnO_4$ _____

40. H_2CO_3 _____

41. H_3PO_3 _____

42. $Se(OH)_2$ _____

Complete and balance the acid-base neutralization reactions. Write "a" under the acid (reactant), and write "b" under the base (reactant). Write "c" under the salt, which is the new ionic compound produced. Hint: carbonate ions (CO_3^{2-}) will split apart to produce CO_2 gas and to help form water.



Section 3 Reinforcement Acids Bases And Salts

A. Whelan



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