

Stoichiometry Worksheet and Key



1. How many moles of O_2 will be formed from 1.65 moles of KClO_3 ?

1.65 mol KClO_3	_____ mol O_2	= _____ mol O_2
	_____ mol KClO_3	

2. How many moles of KClO_3 are needed to make 3.50 moles of KCl ?

3.50 mol KCl		= _____ mol KClO_3

3. How many moles of KCl will be formed from 2.73 moles of KClO_3 ?

_____		= _____



4. How many moles of Fe_2O_3 are produced when 0.275 moles of Fe is reacted?

0.275 mol Fe		= _____ mol Fe_2O_3

5. How many moles of Fe_2O_3 are produced when 31.0 moles of O_2 is reacted?

_____		= _____

6. How many moles of O_2 are needed to react with 8.9 moles of Fe ?

_____		= _____



7. How many moles of O_2 are produced when 1.26 moles of H_2O is reacted?

8. How many moles of H_2O are needed to produce 55.7 moles of H_2 ?

9. If enough H_2O is reacted to produce 3.40 moles of H_2 , then how many moles of O_2 must have been made?
(a bit challenging, but just think about it and you can probably figure it out)

Self Check Reading And Study Workbook Answers

Stoichiometry

Michal Rosen-Zvi



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Resources in Education, 1982-04 **Basic Chemistry** Glenn H. Miller, Frederick B. Augustine, 1975 **Science Books & Films**, 1981 *Bowker's Complete Video Directory*, 1990 *The School World*, 1902 *AAAS Science Book List Supplement* American Association for the Advancement of Science, 1978 Approximately 2700 titles arranged in classified order Each entry gives bibliographical information annotation and reading levels Author and title subject indexes **DAT** Aftab S. Hassan, 1993 *Choice*, 2006 *Science Books*, 1973 *Current Index to Journals in Education*, 1979 *AAAS Science Books*, 1973 *The Software Encyclopedia*, 1988 **Potential Energy Hypersurfaces** Paul G. Mezey, 1987 The importance of the potential surface model has led naturally to a large number of studies on the subject where the emphasis has usually been placed on lower dimensional problems such as the reaction dynamics of diatomic to four atom systems or conformational problems restricted to few internal rotations The purposes and methods of this book are however somewhat different from those of most studies on potential surface problems The emphasis here is placed on those fundamental properties of potential energy hypersurfaces that are general for higher dimensions that is for larger molecules The study of these properties requires some of the tools of global analysis that are not among the routine mathematical techniques of quantum chemists topology homotopy and homology This book provides the reader with an introduction to the fundamentals and to some of the more recent developments in the theory of potential energy hypersurfaces The text is fairly self contained It requires no previous mathematical knowledge from the reader beyond that needed in an undergraduate quantum chemistry course **Resources in Education**, 1982 *Books in Print*, 1993-09 V 1 Authors A D v 2 Authors E K v 3 Authors L R v 4 S Z v 5 Titles A D v 6 Titles E K v 7 Titles L Q v 8 Titles R Z v 9 Out of print out of stock indefinitely v 10 Publishers **The British National Bibliography** Arthur James Wells, 2003 *Paperbound Books in Print*, 1969

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Formulation and Stoichiometry Emil J. Margolis, 2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the serious student this means more often than not the frustrations of questions unanswered The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in contrasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction

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