



COVALENT BONDING

Practice Problems

Do your answers to these very practice problems.

SECTION 8.1 MOLECULAR COMPOUNDS

1. Classify each of the following as ionic or molecular.

- a. NaCl (ionic) b. CH_4 (molecular) c. HCl (molecular)
 d. H_2O (molecular) e. NaOH (ionic)

2. Which of the following is a molecular compound?

- a. H_2O b. NaCl c. NaOH
 d. HCl e. H_2SO_4

3. Which of the following is a molecular compound? (Note: H_2O is a molecular compound.)

a. H_2O (molecular) b. NaCl (ionic) c. HCl (molecular) d. NaOH (ionic) e. H_2SO_4 (molecular)

4. Which of the following is a molecular compound? (Note: H_2O is a molecular compound.)

SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw the Lewis structure for H-F .



2. Draw the Lewis structure for CH_4 .



3. Draw the Lewis structure for CH_3Cl .



4. Draw the Lewis structure for $\text{H-C}\equiv\text{C-H}$.

5. Draw the Lewis structure for H_2O . (Note: H_2O is a molecular compound.)



(Note: H_2O is a molecular compound.)

SECTION 8.3 BONDING: THE ORBITAL

1. Which of the following is a molecular orbital? (Note: H_2O is a molecular orbital.)

2. Which of the following is a molecular orbital? (Note: H_2O is a molecular orbital.)

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CL Gary



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