

SECTION 5.2 ELECTRON ARRANGEMENT IN ATOMS (pages 133–136)

This section shows you how to apply the aufbau principle, the Pauli exclusion principle, and Hund's rule to help you write the electron configurations of elements. It also explains why the electron configurations for some elements differ from those assigned using the aufbau principle.

► Electron Configurations (pages 133–135)

1. The ways in which electrons are arranged around the nuclei of atoms are called _____ electron configurations _____.

Match the name of the rule used to find the electron configurations of atoms with the rule itself.

b _____ 2. aufbau principle

c _____ 3. Pauli exclusion principle

a _____ 4. Hund's rule

- a. When electrons occupy orbitals of equal energy, one electron enters each orbital until all the orbitals contain one electron with parallel spins.
b. Electrons enter orbitals of lowest energy first.
c. An atomic orbital may describe at most two electrons.

5. Look at the aufbau diagram, Figure 5.7 on page 133. Which atomic orbital is of higher energy, a $4f$ or a $5p$ orbital? _____ $4f$

6. Fill in the electron configurations for the elements given in the table. Use the orbital filling diagrams to complete the table.

Element 1	Orbital filling						Electron configuration
	s	2s	2p _x	2p _y	2p _z	3s	
H	☐	☐	☐	☐	☐	☐	1s
He	☐	☐	☐	☐	☐	☐	1s ²
Li	☐	☐	☐	☐	☐	☐	1s ² 2s
C	☐	☐	☐	☐	☐	☐	1s ² 2s ² 2p ²
N	☐	☐	☐	☐	☐	☐	1s ² 2s ² 2p
O	☐	☐	☐	☐	☐	☐	1s ² 2s ² 2p ⁴
F	☐	☐	☐	☐	☐	☐	1s ² 2s ² 2p
Ne	☐	☐	☐	☐	☐	☐	1s ² 2s ² 2p ⁶
Na	☐	☐	☐	☐	☐	☐	1s ² 2s ² 2p ⁶ 3s

Section 4 3 Review Arrangement Of Electrons In Atoms

Ernest Isaac Lewis



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