

CHAPTER 11 REVIEW*Molecular Composition of Gases***MIXED REVIEW**

SHORT ANSWER Answer the following questions in the space provided.

1. _____ The average speed of a gas molecule is most directly related to the _____.
 - (a) polarity of the molecule
 - (b) pressure of the gas
 - (c) temperature of the gas
 - (d) number of moles in the sample
2. _____ Which of the following statements best represents Graham's law for two gases at the same temperature?
 - (a) The molar masses and the velocities are inversely proportional.
 - (b) The molar masses and the velocities are directly proportional.
 - (c) The molar masses and the square roots of the velocities are directly proportional.
 - (d) The molar masses and the squares of the velocities are inversely proportional.
3. Water is synthesized from its elements according to the following equation:
$$2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l})$$
 - _____ a. Which is greater, the volume of H_2 consumed or the volume of O_2 consumed?
 - _____ b. Which is greater, the mass of H_2 consumed or the mass of O_2 consumed?
4. Consider a 2.24 L sample of an ideal gas at STP.
 - _____ a. How many moles are present in the sample, or is more data needed?
 - _____ b. How many molecules are present in the sample, or is more data needed?
 - _____ c. What is the mass of the sample, or is more data needed?

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

5. _____ Calculate the volume of 3.5 mol of an ideal gas at 24°C and 760 torr.

Section 3 Molecular Composition Of Gases Answers

John T. Moore, Richard H. Langley



Section 3 Molecular Composition Of Gases Answers:

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Chemistry for the Utterly Confused John T. Moore,Richard H. Langley,2007-06-01

Banish bafflement in this tough subject From formulas and lab techniques to the periodic table

Chemistry for the Utterly Confused

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