

Chapter 1: Answers to Review Problems

Section 1.1

- 1) The energy of electromagnetic radiation is inversely proportional to its wavelength; the shorter the wavelength, the higher the energy.
The frequency of a wave refers to the number of times a wave repeats itself in a given amount of time. The shorter the wavelength, the higher the frequency.
Energy and frequency are directly proportional; the higher the frequency, the higher the energy.
- 2) Each of these series refers to a group of spectral lines, each line corresponding to a specific wavelength of electromagnetic radiation, and therefore, a specific energy of radiation.
The Bohr model states that normally the electron in hydrogen occupies the lowest energy level and this is called the ground state. If the electron absorbs energy, it is promoted to a higher energy level, known as an excited state. As the excited state is not stable, the electron will eventually return to a lower energy level, emitting the absorbed energy as light. The energy difference between the excited state and the ground state corresponds to light of a certain wavelength.
The Balmer series refers to wavelengths of light emitted when an electron falls from a higher energy level to $n = 2$; these wavelengths fall in the visible region of the spectrum.
The Lyman series refers to wavelengths of light emitted when an electron falls from a higher energy level to $n = 1$; these wavelengths fall in the ultraviolet region of the spectrum.
The Paschen series refers to wavelengths of light emitted when an electron falls from a higher energy level to $n = 3$; these wavelengths fall in the infrared region of the spectrum.
- 3) The $\Psi_2(x)$ wave has a shorter wavelength than that of the $\Psi_1(x)$ wave. Therefore, the energy of the $\Psi_2(x)$ wave is greater than that of the $\Psi_1(x)$ wave. Also, the greater the number of nodes, the higher the energy.
- 4) **Pauli's Exclusion Principle** states that no two electrons in an atom will have the same set of four quantum numbers. The valence shell configuration of Carbon is $2s^2 2p^2$
For the two electrons in the $2s$:
Electron # 1: $n = 2, l = 0, m_l = 0, m_s = +\frac{1}{2}$
Electron # 2: $n = 2, l = 0, m_l = 0, m_s = -\frac{1}{2}$
These two electrons have different values for m_s , so the two electrons have a different set of 4 quantum numbers.

Section Review Chapter 1 answers

Rosina Ehmann



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Table of Contents Section Review Chapter 1answers

1. Understanding the eBook Section Review Chapter 1answers
 - The Rise of Digital Reading Section Review Chapter 1answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Section Review Chapter 1answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section Review Chapter 1answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section Review Chapter 1answers
 - Personalized Recommendations
 - Section Review Chapter 1answers User Reviews and Ratings
 - Section Review Chapter 1answers and Bestseller Lists
5. Accessing Section Review Chapter 1answers Free and Paid eBooks
 - Section Review Chapter 1answers Public Domain eBooks
 - Section Review Chapter 1answers eBook Subscription Services
 - Section Review Chapter 1answers Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Section Review Chapter 1 answers Compatibility with Devices
 - Section Review Chapter 1 answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section Review Chapter 1 answers
 - Highlighting and Note-Taking Section Review Chapter 1 answers
 - Interactive Elements Section Review Chapter 1 answers
8. Staying Engaged with Section Review Chapter 1 answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section Review Chapter 1 answers
9. Balancing eBooks and Physical Books Section Review Chapter 1 answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section Review Chapter 1 answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Section Review Chapter 1 answers
 - Setting Reading Goals Section Review Chapter 1 answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Section Review Chapter 1 answers
 - Fact-Checking eBook Content of Section Review Chapter 1 answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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