

Stoichiometry

Quiz 11*9 Section 9

Section 9: Introduction to Stoichiometry

In the space provided, write the letter of the best response that best completes each problem or best answers each question.

1. Stoichiometry is the branch of chemistry that deals with, according to convention, and with the basic relationships between:
 - a. elements
 - b. compounds
 - c. mass relationships
 - d. physical measurements
2. Some basic laws governing all chemical reactions are:
 - a. the law of conservation of mass.
 - b. the law of gas laws.
 - c. the law of conservation of momentum.
 - d. temperature control.
3. The number of significant figures in a product or quotient is usually the same as the number of significant figures in:
 - a. the number of significant figures in the factor with the least number of significant figures.
 - b. the number of significant figures in the factor with the most significant figures.
 - c. the number of significant figures of any component given in the problem.
 - d. the number of significant figures in the sum of the numbers in the chemical equation.
4. The chemical equation $\text{Fe}_2\text{S}_3 + 10\text{H}_2\text{O} \rightarrow 2\text{Fe}(\text{OH})_3 + 3\text{H}_2\text{S}$ can be interpreted correctly as:
 - a. 1 molecule of Fe_2S_3 + 10 molecules of H_2O form 2 molecules of $\text{Fe}(\text{OH})_3$.
 - b. 1 mole of Fe_2S_3 + 10 moles of H_2O form 2 moles of $\text{Fe}(\text{OH})_3$.
 - c. 1 g of Fe_2S_3 + 10 g of H_2O form 2 g of $\text{Fe}(\text{OH})_3$.
 - d. both (a) and (b).
5. How many moles of $\text{Fe}(\text{OH})_3$ are formed from the following chemical equation?
 - a. 1
 - b. 2
 - c. 3
 - d. 4



Section Quiz Introduction To Stoichiometry

G Thomas



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