

GUIDED PRACTICE PROBLEM 34 (page 307)

34. Calculate the percent composition of these compounds.

- a. ethane (C_2H_6)
 b. sodium bisulfate (NaHSO_4)

a. ethane (C_2H_6)**Analyze****Step 1.** List the knowns and the unknowns.**Knowns**mass of C in one mole ethane = $2 \times 12.0 \text{ g} = 24.0 \text{ g}$ mass of H in one mole ethane = $6 \times 1.0 \text{ g} = 6.0 \text{ g}$ molar mass of $\text{C}_2\text{H}_6 = 24.0 \text{ g} + 6.0 \text{ g} = 30.0 \text{ g}$ **Unknowns**

percent C = ? %

percent H = ? %

Because no masses are given, the percent composition can be determined based on the molar mass of the substance. The percent of an element in a compound is the mass of the element in the compound divided by the mass of the compound. To express the ratio as a percent, the ratio is multiplied by 100%.

Calculate**Step 2.** Solve for the unknown.

$$\text{percent C} = \frac{24.0 \text{ g C}}{30.0 \text{ g compound}} \times 100\% = 80.0\% \text{ C}$$

$$\text{percent H} = \frac{6.0 \text{ g H}}{30.0 \text{ g compound}} \times 100\% = 20.0\% \text{ H}$$

Evaluate**Step 3.** Does the result make sense?

The percents of the elements of the compound add up to 100%.

$$80.0\% + 20.0\% = 100\%$$

Section Chemical Quantities Practice Problem Answers

**Peter J. Mikulecky, Katherine
Brutlag, Michelle Rose Gilman, Brian
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Table of Contents Section Chemical Quantities Practice Problem Answers

1. Understanding the eBook Section Chemical Quantities Practice Problem Answers
 - The Rise of Digital Reading Section Chemical Quantities Practice Problem Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Section Chemical Quantities Practice Problem Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section Chemical Quantities Practice Problem Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section Chemical Quantities Practice Problem Answers
 - Personalized Recommendations
 - Section Chemical Quantities Practice Problem Answers User Reviews and Ratings
 - Section Chemical Quantities Practice Problem Answers and Bestseller Lists
5. Accessing Section Chemical Quantities Practice Problem Answers Free and Paid eBooks

- Section Chemical Quantities Practice Problem Answers Public Domain eBooks
- Section Chemical Quantities Practice Problem Answers eBook Subscription Services
- Section Chemical Quantities Practice Problem Answers Budget-Friendly Options
- 6. Navigating Section Chemical Quantities Practice Problem Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Section Chemical Quantities Practice Problem Answers Compatibility with Devices
 - Section Chemical Quantities Practice Problem Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section Chemical Quantities Practice Problem Answers
 - Highlighting and Note-Taking Section Chemical Quantities Practice Problem Answers
 - Interactive Elements Section Chemical Quantities Practice Problem Answers
- 8. Staying Engaged with Section Chemical Quantities Practice Problem Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section Chemical Quantities Practice Problem Answers
- 9. Balancing eBooks and Physical Books Section Chemical Quantities Practice Problem Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section Chemical Quantities Practice Problem Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Section Chemical Quantities Practice Problem Answers
 - Setting Reading Goals Section Chemical Quantities Practice Problem Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section Chemical Quantities Practice Problem Answers
 - Fact-Checking eBook Content of Section Chemical Quantities Practice Problem Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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