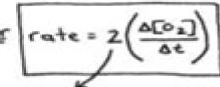
- Reaction Rate = the change in concentration of 12-3 reactant(s) or products per unit time.

in the reaction on the last page (Table 12.1; Fig. 12.1) we have:

because being 2 NO2(9) -> 2 NO(9) + O2(9)

$$rate = -\frac{\Delta[NO_L]}{\Delta t}$$



because only half as much produced. See the balanced equation.

* so, the <u>reaction rate</u>, or "<u>rate</u>", can be written in terms of any of the reactants or in terms of any of the products.

To be consistent however, it's customary to represent the reaction rate in terms of the reactant (s), so: * [coto = - A[Noz]

- An Examination of Figure 12-1 on the Previous Page.

calculate the instantaneous rate of the reaction.

Ly = calculated by finding the slope (m) of the line tangent to the curve at the specific time (t).

AW Rasmussen

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