

Wintersemester 2019/2020 - Klausuraufgaben 8B

Aufgabe 8 (max. 70 Minuten/100 Punkte)

Die nachstehenden Probleme sollen Sie zur Vertiefung des wissenschaftlichen Verständnisses der Stoffeigenschaften (Kapitel 10) befähigen. Sie sind, soweit in der Aufgabenstellung nicht anders angegeben, unter folgenden Bedingungen zu lösen:

$$p = p_0 = 1 \text{ bar} = 10^5 \text{ Pa}$$

1. Chloroform (Molare Masse $M = 119,38 \text{ g/mol}$) wird bei einer Temperatur von 25°C verdunstet. Wie groß ist die molare Verdampfungsenthalpie?

$$\Delta H_{\text{V}} = \frac{119,38 \text{ g/mol} \cdot 10^3 \text{ J/kg} \cdot 25^\circ\text{C}}{10^5 \text{ Pa} \cdot 10^{-3} \text{ m}^3/\text{mol}} = 29,85 \text{ kJ/mol}$$

2. Ethylacetat (Molare Masse $M = 88,10 \text{ g/mol}$) wird bei 25°C verdunstet. Wie groß ist die molare Verdampfungsenthalpie bei 25°C ?

$$\Delta H_{\text{V}} = \frac{88,10 \text{ g/mol} \cdot 10^3 \text{ J/kg} \cdot 25^\circ\text{C}}{10^5 \text{ Pa} \cdot 10^{-3} \text{ m}^3/\text{mol}} = 22,03 \text{ kJ/mol}$$

3. Ethanol ist ein verdampfendes flüchtiges Gas. Bei einer Temperatur von 25°C beträgt die molare Verdampfungsenthalpie $38,56 \text{ kJ/mol}$. Wie groß ist die molare Verdampfungsenthalpie bei 50°C ?

$$\Delta H_{\text{V}} = \frac{38,56 \text{ kJ/mol} \cdot 10^3 \text{ J/kJ} \cdot 25^\circ\text{C}}{10^5 \text{ Pa} \cdot 10^{-3} \text{ m}^3/\text{mol}} = 9,64 \text{ kJ/mol}$$

4. Die molare Verdampfungsenthalpie von Ethanol beträgt $38,56 \text{ kJ/mol}$. Wie groß ist die molare Verdampfungsenthalpie von Ethanol bei 50°C ? Wie groß ist die molare Verdampfungsenthalpie von Ethanol bei 75°C ? Wie groß ist die molare Verdampfungsenthalpie von Ethanol bei 100°C ? Wie groß ist die molare Verdampfungsenthalpie von Ethanol bei 125°C ? Wie groß ist die molare Verdampfungsenthalpie von Ethanol bei 150°C ?

$$\Delta H_{\text{V}} = \frac{38,56 \text{ kJ/mol} \cdot 10^3 \text{ J/kJ} \cdot 25^\circ\text{C}}{10^5 \text{ Pa} \cdot 10^{-3} \text{ m}^3/\text{mol}} = 9,64 \text{ kJ/mol}$$

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6. Ethanol ist ein verdampfendes flüchtiges Gas. Bei einer Temperatur von 25°C beträgt die molare Verdampfungsenthalpie $38,56 \text{ kJ/mol}$. Wie groß ist die molare Verdampfungsenthalpie von Ethanol bei 50°C ? Wie groß ist die molare Verdampfungsenthalpie von Ethanol bei 75°C ? Wie groß ist die molare Verdampfungsenthalpie von Ethanol bei 100°C ? Wie groß ist die molare Verdampfungsenthalpie von Ethanol bei 125°C ? Wie groß ist die molare Verdampfungsenthalpie von Ethanol bei 150°C ?

Section 3 Gases Ch 1review Answers

JG Myers



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