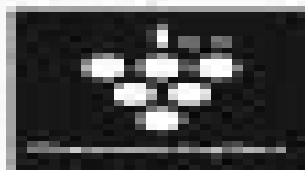
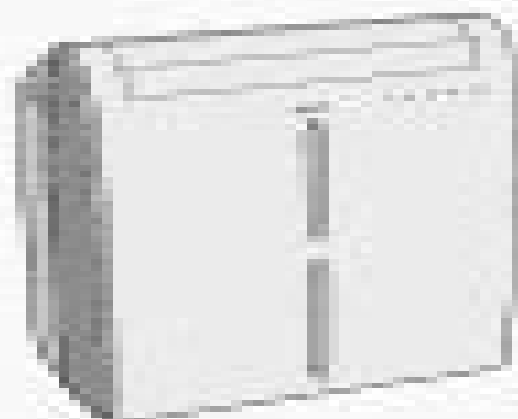


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# Sharp Cv P10mc Manual

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fraction  $x$  is the ratio of moles of one substance in a mixture to the total number of moles of all substances for a mixture of two substances  $c_e a$  and  $c_e b$  the mole fractions of each would be written as

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mass  $C_6H_6$  6 12 6 1 g mol 1 78 g mol 1 hence  $C_6H_6$  s number of moles  $30.78 \text{ mol} / 384.6 \text{ g mol}^{-1}$  molar mass of carbon tet  $CCl_4$  12 4 355 154 g mol 1

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web since the mole fraction of the  $O_2$  is 0.050 it contributes 0.050 or if you will 5.0% of the total pressure thus total pressure is 0.2105042 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of  $O_2$  are mixed in a flask at room temperature

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