

Periodic Trends Worksheet

Directions: Use the table to determine the following properties:

1. Write the following elements by increasing atomic radius: sodium, sulfur, aluminum, oxygen, potassium.

K, Na, Al, S, O

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, aluminum, potassium.

O, S, Al, Na, K

3. Which atom has the highest first ionization energy than sodium?

None, potassium is the lowest first ionization energy and sodium is the highest.

4. Which two elements in the same family generally have similar properties?

There have been two atoms that behave a little similar with potassium.

5. Evaluate whether the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius (increases to the right) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms of the periodic table? Write numbers, this period?

Periods 1 through 7, including lanthanide and actinide series.

7. Which period in the periodic table shows atoms of the periodic table? Which period shows the longest?

Period 7, including lanthanide and actinide series.

8. Write the atomic number and the first ionization energy.

1	H	1
2	He	2
3	Li	3
4	Be	4
5	B	5
6	C	6
7	N	7
8	O	8
9	F	9
10	Ne	10

11	Na	11
12	Mg	12
13	Al	13
14	Si	14
15	P	15
16	S	16
17	Cl	17
18	Ar	18

9. Write the atomic number and the first ionization energy.

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